

Chemistry 534 Test: Electrochemistry
(ref:test:electrochem.)

Name:
Date:

Matching questions: Place the correct letter in front of the corresponding term. (2 points)

- A) a tube containing a conducting solution
- B) an electrochemical cell that is used to convert chemical energy to electrical energy
- C) a voltaic cell in which the electrolyte is a paste
- D) a conductor in a circuit that carries electrons to or from a substance other than a metal
- E) a process in which electrical energy is used to bring about a chemical change

_____ dry cell; _____ electrode; _____ electrolysis; _____ salt bridge; _____ voltaic cell;

Multiple-Choice: Select the best answer, either (A, B, C, or D). Completely darken the letter in the answer section at the end. (30 points)

1. What happens in an electrochemical cell?
 - A) Helium gas is produced.
 - B) Mechanical and light energy are interconverted.
 - C) Chemical and electrical energy are interconverted.
 - D) Chemistry tests are shared and human interactions are simulated.
2. How can a redox reaction be used as a source of electrical energy?
 - A) Two half reactions must be physically separated.
 - B) One half-reaction must involve two metals.
 - C) Two half-reactions must involve more than one electron.
 - D) One half-reaction must use a wire electrode.
3. Ions of which of the following metals are the most easily reduced?
 - A) potassium
 - B) calcium
 - C) iron
 - D) lead
 - E) mercury
4. A zinc-copper cell is constructed: $Zn | Zn^{2+} (1M) || Cu^{2+} | Cu$
 What occurs to the mass of the copper electrode as the reaction proceeds? (Zinc is above copper in the activity series of metals.)
 - A) it decreases
 - B) it increases
 - C) it remains the same

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Katrin Zwirgmaier



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