

Chapter 9 – Stoichiometry – Chapter Review**Last page worked out for you.**

Match each description on the right to the correct term on the left.

- | | | |
|-----------------|--|----------------------|
| <u>B</u> | 1. The substance that determines the amount of product that can be formed in a reaction. | A. Percent yield |
| <u>E</u> | 2. The amount of product that forms when a reaction is carried out in the laboratory. | B. Limiting reagent |
| <u>D</u> | 3. The calculation of quantities in chemical equations. | C. Theoretical yield |
| <u>A</u> | 4. The ratio of the actual yield to the theoretical yield expressed as a percent. | D. Stoichiometry |
| <u>F</u> | 5. The substance that is present in enough quantity to react with a limiting reagent. | E. Actual yield |
| <u>C</u> | 6. The maximum amount of products that could be formed from given amounts of reactants. | F. Excess reagent |

Multiple Choice - Choose the best answer and write its letter in the blank.

- D** 7. How many moles of H₂O would be required to produce 2.5 moles of O₂?
- $$2 \text{H}_2\text{O} \rightarrow 2 \text{H}_2 + \text{O}_2$$
- A. 2.0 mol C. 4.0 mol
B. 2.5 mol D. 5.0 mol
- D** 8. If 3.00 mol of CaCO₃ undergo decomposition to form CaO and CO₂, how many grams of CO₂ are produced?
- $$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$$
- A. 3.00 g C. 88.0 g
B. 44.0 g D. 132 g
- B** 9. How many moles of H₂O are produced when 240 grams of CuO react?
- $$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$$
- A. 1.0 mol C. 18 mol
B. 3.0 mol D. 54 mol

Chapter Review Stoichiometry Answer Key

Stanley R. Radel, Marjorie H. Navidi



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