

Chemistry MSL Review

Math in Chemistry

1. Give a definition and example of a qualitative measurement. Give a definition and example of a quantitative measurement.
2. How many sig figs are contained in the following numbers:
 0.067 2450 2450.0 0.0100
3. Define accuracy and precision.
4. If a block has a known density of 1.45, what is the percent error of the value 1.14 measured in lab?
5. A block of unknown composition has a mass of 23.2 grams. Using water displacement, its volume is measured at 27.9 mL. What is the block's density?
6. What is the volume of an object with a density of 1.94 g/mL and a mass of 15.5g?
7. Convert the temperature 88°C to Kelvin.
8. Metric conversions: Convert 4.58 m to cm. Convert 104.5 J to kJ. Convert 0.009 332 g to mg.
9. What is the difference between a chemical change and a physical change?

Atomic Structure

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.



1. Define the following terms:
 atomic number; mass number; isotope; atomic mass
2. How do you determine the number of protons in an atom? How do you determine the number of neutrons in an atom?
3. Complete the following table:

Isotope	Atomic #	# Protons	# Neutrons	Mass #
$^{12}_6\text{C}$				
	17		15	
		26		56
		17		37
			2	3
	52			128
		50	70	

- a. Consider the second and fourth row in the table above. What do they have in common?
- b. Consider the second and fourth row in the table above. What are the differences?

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Ying-Ying Zheng



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Table of Contents Dpi Study Guide Msl Chemistry

1. Understanding the eBook Dpi Study Guide Msl Chemistry
 - The Rise of Digital Reading Dpi Study Guide Msl Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Dpi Study Guide Msl Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Dpi Study Guide Msl Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Dpi Study Guide Msl Chemistry
 - Personalized Recommendations
 - Dpi Study Guide Msl Chemistry User Reviews and Ratings
 - Dpi Study Guide Msl Chemistry and Bestseller Lists
5. Accessing Dpi Study Guide Msl Chemistry Free and Paid eBooks
 - Dpi Study Guide Msl Chemistry Public Domain eBooks
 - Dpi Study Guide Msl Chemistry eBook Subscription Services
 - Dpi Study Guide Msl Chemistry Budget-Friendly Options

6. Navigating Dpi Study Guide Msl Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Dpi Study Guide Msl Chemistry Compatibility with Devices
 - Dpi Study Guide Msl Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Dpi Study Guide Msl Chemistry
 - Highlighting and Note-Taking Dpi Study Guide Msl Chemistry
 - Interactive Elements Dpi Study Guide Msl Chemistry
8. Staying Engaged with Dpi Study Guide Msl Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Dpi Study Guide Msl Chemistry
9. Balancing eBooks and Physical Books Dpi Study Guide Msl Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Dpi Study Guide Msl Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Dpi Study Guide Msl Chemistry
 - Setting Reading Goals Dpi Study Guide Msl Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Dpi Study Guide Msl Chemistry
 - Fact-Checking eBook Content of Dpi Study Guide Msl Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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