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# Teacher's resource answers

## Chapter 21

1  $K_a = 10^{-7.5} = 10^{-7.5}$  [1]

$= 6.31 \times 10^{-8} \text{ mol dm}^{-3}$  [2]

(take 1 mark off if units are incorrect)

$[\text{H}^+] \sqrt{K_a \times [\text{H}_2\text{PO}_4^-]} = 7.94 \times 10^{-5} \text{ mol dm}^{-3}$

(this is because  $K_a = \frac{[\text{H}^+][\text{HPO}_4^{2-}]}{[\text{H}_2\text{PO}_4^-]}$ ) [1]

$\text{pH} = -\log_{10}[\text{H}^+] = 4.10$  [1]

2 a A buffer is a mixture of a weak acid and its conjugate base. [1]

It resists changes in pH. [1]

b The equilibrium present in the buffer is  $\text{H}_2\text{PO}_4^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}^+$  [1]

If unwanted  $\text{H}^+$  ions are added they combine with the  $\text{HPO}_4^{2-}$  [1]

to form the largely undissociated weak acid  $\text{H}_2\text{PO}_4^-$  and are therefore removed, leaving the pH near constant. [1]

If unwanted  $\text{OH}^-$  ions are added they are neutralised by the  $\text{H}^+$  ions thus disturbing the equilibrium. [1]

More  $\text{H}_2\text{PO}_4^-$  dissociates until the  $K_a$  returns to its proper value and pH remains near constant. [1]

c  $K_a = \frac{[\text{H}^+][\text{HPO}_4^{2-}]}{[\text{H}_2\text{PO}_4^-]}$  [1]

rearranging the formula gives

$[\text{H}^+] = K_a \times \frac{[\text{H}_2\text{PO}_4^-]}{[\text{HPO}_4^{2-}]}$  [1]

$[\text{H}^+] = 6.31 \times 10^{-8} \times \frac{0.01}{0.06} = 1.05 \times 10^{-8} \text{ mol dm}^{-3}$  [1]

$\text{pH} = -\log_{10}[\text{H}^+] = 7.98$  [1]

d The equilibrium  $\text{H}_2\text{PO}_4^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}^+$  would be displaced to the left [1]

in order to remove the added  $\text{H}_2\text{PO}_4^-$  ions (Le Chatelier's principle). [1]

Therefore,  $[\text{H}^+]$  increases and pH decreases. [1]

3 a  $\text{Ca}^{2+}$  and  $\text{PO}_4^{3-}$  [2]

b  $\text{Ca}_3(\text{PO}_4)_2 \rightleftharpoons 3\text{Ca}^{2+} + 2\text{PO}_4^{3-}$  [1]

c  $K_{sp} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2$  [1]

d solubility =  $2.48 \times 10^{-8} \text{ mol dm}^{-3}$  [1]

$[\text{Ca}^{2+}] = 3 \times 2.48 \times 10^{-8} = 7.44 \times 10^{-8} \text{ mol dm}^{-3}$  [1]

and  $[\text{PO}_4^{3-}] = 2 \times 2.48 \times 10^{-8} = 4.96 \times 10^{-8} \text{ mol dm}^{-3}$  [1]

$K_{sp} = (7.44 \times 10^{-8})^3 \times (4.96 \times 10^{-8})^2 = 1.01 \times 10^{-28} \text{ mol}^5 \text{ dm}^{-15}$  [2]

(take one mark off for incorrect units)

e There is a strong attraction between the +2 calcium ions and the -3 phosphate ions. [1]

Therefore, when the gaseous ions form the ionic lattice a lot of energy is produced as strong ionic bonds are formed. [1]

f The enthalpies of hydration of the  $\text{Ca}^{2+}$  ions and the  $\text{PO}_4^{3-}$  ions are not sufficiently negative (enough) [1]

to overcome the large amount of energy needed to produce the gaseous ions from the ionic lattice (-L.E.). [1]

# Chapter Resource Answers

**Leslie A. Swetnam**



## Chapter Resource Answers:

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*Instructor's Resource Manual with Test Items: Assessment* John Salvia,1991

*Louisiana: The Land and Its People (Teacher's Resource Book)* Sue Eakin,Mrs Manie Culbertson,1998-11-30 Louisiana Culture Comes Alive An in depth discussion of Louisiana and the forces that have shaped it characterize Louisiana The Land and Its People Available from Pelican Publishing Company the 560 page volume provides a comprehensive overview of Louisiana history from the first Indian inhabitants through the Civil War and Reconstruction to modern times Some 200 photographs and illustrations including 45 maps complement the text and encourage the reader to delve further into the background of Louisiana and its people Ideal for classroom use a teacher s guide is also available Louisiana The Land and Its People is much more than a textbook Its detailed discussion of the factors that have molded the state provides insights into current issues and solutions Everyone interested in understanding Louisiana both today and yesterday will benefit from reading this book

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