

Covalent Bonding Practice Worksheet

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705

Explanation: Covalent compounds are compounds between nonmetal atoms. The atoms "start" with a number of electrons equal to their group number in the periodic table. Then they share electrons so that all the atoms have filled shells with a "magic number" of electrons. The magic number is 2 for hydrogen and 8 for C, N, O, F, Si, P, S, Cl, Se, Br and I. Remember that compounds made with a metal atom and a nonmetal atom make an ionic compound, not a covalent one, and compounds of metals are also different, these have metallic bonds and do not require a strict ratio.

Instructions: Fill in the blanks in the table. The first row is done for you; study the first row carefully before you start, especially the difference between different boxes (remember there are many ways to draw the same compound). Some rows have hints. If a box has an X you don't have to fill in that box.

Helpful Tips: 1) All atoms start with a number of outer shell electrons = their group number. You cannot change the element's group number. 2) If you have to choose between two shapes, the symmetrical, balanced one is usually best. 3) Try to figure the structures out yourself, don't cheat and use the Internet. ☺

Compound Formula	Full electronic structure	Lewis structure	Dot-cross diagram	Structural formula	3-D formula
H ₂ O		H:O:H ••	H:O:H ••	H-O-H	
NH ₃		H:N:H •• H	H:N:H •• H	H-N-H H	
CH ₃ Br		H:C:H •• Br	H:C:H •• Br	H-C-H Br	

pts each
(12)
(12)

Chapter Covalent Bonding Practice Problems Answers

David R. Klein, Laurie S. Starkey



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Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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Organic Chemistry David R. Klein, Laurie S. Starkey, 2024-12-27 In the 5th Edition of Organic Chemistry David Klein continues to set the standard for how students learn by building on his innovative SkillBuilder approach enabling learners to effectively grasp the complex language of organic chemistry through structured guided practice Joining David Klein for this edition as an author is longtime collaborator Laurie Starkey Cal Poly Pomona whose classroom creativity digital expertise and positive teaching style bring a fresh perspective to Organic Chemistry Her contributions enhance the proven SkillBuilder method infusing it with new pedagogically relevant photo examples that make the material even more accessible and engaging for students The new edition is thoughtfully updated with extensive content revisions refined SkillBuilders and fresh examples all shaped by valuable feedback from instructors It also introduces a wider range of diverse examples vivid illustrations and practical applications tailored to both Organic Chemistry I and II Together Klein and Starkey have crafted a comprehensive and dynamic resource that blends proven techniques with fresh insights ensuring the best learning experience for students

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