

Basic Atomic Structure Key

Complete the following using your Periodic Table as reference.

1. 19 Atomic number of potassium
2. 30.97 Mass number of phosphorus
3. 7 Protons in nitrogen
4. 6 Electrons in carbon
5. 12 Neutrons in sodium
6. 13 Atomic number of aluminum
7. 40.08 Mass number of calcium
8. 18 Protons in argon
9. 23 Electrons in vanadium
10. 24 Neutrons in scandium
11. 61 Atomic number of promethium
12. 152.0 Mass number of europium
13. 65 Protons in terbium
14. 47 Electrons in silver
15. 78 Neutrons in cesium
16. 35 Atomic number of bromine
17. 186.2 Mass number of rhenium
18. 77 Protons in iridium
19. 78 Electrons in platinum
20. 118 Neutrons in gold

Isotopes

1. There are 5 naturally occurring isotopes of zinc: . None of these isotopes has the atomic mass of 65.41, as listed on the periodic table for zinc. Explain why.

The mass number listed on the periodic table is an average abundance of all the isotopes.

2. Two isotopes of gallium occur naturally, . Which isotope is more abundant? How do you know?

The mass number of gallium on the periodic table is listed as 69.72, which is closer to 69 than 71, which indicates that it is the more abundant isotope.

Chapter Atomic Structure Wordwise Answers

Edward Uhler Condon



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