

Chapter Test B

Chapter: Stoichiometry

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. Knowing the mole ratio of a reactant and product in a chemical reaction would allow you to determine
- the energy released in the reaction.
 - the speed of the reaction.
 - the mass of the product produced from a known mass of reactant.
 - whether the reaction is reversible.
- _____ 2. In the equation $6\text{CO}_2 + 6\text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2$, the mole ratio of water to oxygen is
- 2:1.
 - 1:1.
 - 3:2.
 - 6:8.
- _____ 3. In the equation $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$, the mole ratio of calcium hydroxide to aluminum hydroxide is
- 1:3.
 - 2:3.
 - 1:1.
 - 3:2.
- _____ 4. For the reaction represented by the equation $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, how many moles of nitrogen are required to produce 18 mol of ammonia?
- 9
 - 18
 - 27
 - 36
- _____ 5. In the reaction represented by the equation $\text{A} + \text{B} \rightarrow \text{C} + \text{D}$, if there is more reactant B than is required to completely react with all of A, then
- A is the limiting reactant.
 - B is the limiting reactant.
 - there is no limiting reactant.
 - no product can be formed.

Chapter 9 Stoichiometry

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Fuller Holtzclaw, William R. Robinson**

Chapter 9 Stoichiometry:

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