

CHAPTER 9 REVIEW

Stoichiometry**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- _____ The coefficients in a chemical equation represent the
 - masses in grams of all reactants and products.
 - relative number of moles of reactants and products.
 - number of atoms of each element in each compound in a reaction.
 - number of valence electrons involved in a reaction.
- _____ Which of the following would not be studied within the topic of stoichiometry?
 - the mole ratio of Al to Cl in the compound aluminum chloride
 - the mass of carbon produced when a known mass of sucrose decomposes
 - the number of moles of hydrogen that will react with a known quantity of oxygen
 - the amount of energy required to break the ionic bonds in CaF_2
- _____ A balanced chemical equation allows you to determine the
 - mole ratio of any two substances in the reaction.
 - energy released in the reaction.
 - electron configuration of all elements in the reaction.
 - reaction mechanism involved in the reaction.
- _____ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
 - reaction sequence.
 - bond energy.
 - mole ratio.
 - element proportion.
- Given the reaction represented by the following unbalanced equation:
 $\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$
 - Balance the equation.

b. What is the mole ratio of NO_2 to O_2 ?

c. If 20.0 mol of NO_2 form, how many moles of O_2 must have been consumed?

d. Twice as many moles of NO_2 form as moles of N_2O are consumed. True or False?

e. Twice as many grams of NO_2 form as grams of N_2O are consumed. True or False?

Chapter 9 Section Stoichiometry Answers

Robert Nakon



Chapter 9 Section Stoichiometry Answers:

Programmed Problem-solving for First-year Chemistry Arnold B. Loebel,1983 *Chemical Problem-solving by Dimensional Analysis* Arnold B. Loebel,1978 *EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS* CHANG,2013-01-07 *EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS* **McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition** Thomas Evangelist,2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill s SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies *McGraw-Hill's SAT Subject Test: Chemistry, 2ed* Thomas Evangelist,2009-02-01 We want to help you score high on the SAT Chemistry test We ve put all of our proven expertise into McGraw Hill s SAT Subject Test Chemistry to make sure you re fully prepared for this difficult exam With this book you ll get essential skill building techniques and strategies created by a leading high school chemistry teacher You ll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill s SAT Subject Test Chemistry we ll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score [Instructor's Manual and Test Bank to Accompany Basic Concepts of Chemistry](#) Malone,2000-04-18 [Physiological Ecology](#) William H. Karasov,Carlos Martínez del Río,2020-05-05 Unlocking the puzzle of how animals behave and how they interact with their environments is impossible without understanding the physiological processes that determine their use of food resources But long overdue is a user friendly introduction to the subject that systematically bridges the gap between physiology and ecology Ecologists for whom such knowledge can help clarify the consequences of global climate change the biodiversity crisis and pollution often find themselves wading through an unwieldy technically top heavy literature Here William Karasov and Carlos Martínez del Río present the first accessible and authoritative one volume overview of the physiological and biochemical principles that shape how animals procure energy and nutrients and free themselves of toxins and how this relates to broader ecological phenomena After introducing primary concepts the authors review the chemical ecology of food and then discuss how animals digest and process food Their broad view includes symbioses and extends even to ecosystem phenomena such as

ecological stoichiometry and toxicant biomagnification They introduce key methods and illustrate principles with wide ranging vertebrate and invertebrate examples Uniquely they also link the physiological mechanisms of resource use with ecological phenomena such as how and why animals choose what they eat and how they participate in the exchange of energy and materials in their biological communities Thoroughly up to date and pointing the way to future research

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