

Section 18.4 Neutralization

In your textbook, read about neutralization and titration.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- _____ 1. A chemical dye that changes color based on the pH of a solution
- _____ 2. A method for using a neutralization reaction to determine the concentration of a solution
- _____ 3. A reaction in which an acid and a base react to produce a salt and water
- _____ 4. A solution of known concentration
- _____ 5. An ionic product of an acid-base reaction
- _____ 6. The point in a titration in which an indicator changes color
- _____ 7. The stoichiometric point of a titration

Column B

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. salt
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

Acid	Base	Salt formula	Salt name
8. HCl	KOH	KCl	potassium chloride
9. H_2SO_4	$Mg(OH)_2$		
10. H_3PO_4	NaOH		
11. HNO_3	$Fe(OH)_3$		
12. H_3PO_4	$Ca(OH)_2$		

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

Sequence of Steps

- _____ 13. Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- _____ 14. Fill a buret with the standard solution.
- _____ 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

Chapter 9 Review Stoichiometry Key

**Allan Blackman, Daniel
Southam, Gwendolyn Lawrie, Natalie
Williamson, Christopher Thompson**

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Table of Contents Chapter 9 Review Stoichiometry Key

1. Understanding the eBook Chapter 9 Review Stoichiometry Key
 - The Rise of Digital Reading Chapter 9 Review Stoichiometry Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Review Stoichiometry Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Review Stoichiometry Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Review Stoichiometry Key
 - Personalized Recommendations
 - Chapter 9 Review Stoichiometry Key User Reviews and Ratings
 - Chapter 9 Review Stoichiometry Key and Bestseller Lists
5. Accessing Chapter 9 Review Stoichiometry Key Free and Paid eBooks
 - Chapter 9 Review Stoichiometry Key Public Domain eBooks
 - Chapter 9 Review Stoichiometry Key eBook Subscription Services
 - Chapter 9 Review Stoichiometry Key Budget-Friendly Options

6. Navigating Chapter 9 Review Stoichiometry Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Review Stoichiometry Key Compatibility with Devices
 - Chapter 9 Review Stoichiometry Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Review Stoichiometry Key
 - Highlighting and Note-Taking Chapter 9 Review Stoichiometry Key
 - Interactive Elements Chapter 9 Review Stoichiometry Key
8. Staying Engaged with Chapter 9 Review Stoichiometry Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Review Stoichiometry Key
9. Balancing eBooks and Physical Books Chapter 9 Review Stoichiometry Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Review Stoichiometry Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 9 Review Stoichiometry Key
 - Setting Reading Goals Chapter 9 Review Stoichiometry Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 9 Review Stoichiometry Key
 - Fact-Checking eBook Content of Chapter 9 Review Stoichiometry Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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