

## CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfide, the roman numeral tells us (2)
  - How many atoms of Fe are in one formula unit.
  - How many sulfide ions can be obtained by the ion atom.
  - The charge on each Fe ion.
  - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is (1)
  - Change the number of ions represented in the formula.
  - Change the charges on the other ions in the compound.
  - Change the formula so that it no longer represents the compound.
  - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{CH}_3\text{CO}_2\text{H}$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead II carbonate	$\text{PbCO}_3$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Sulfuric acid	$\text{H}_2\text{SO}_4$
Nitrogen pentoxide	$\text{N}_2\text{O}_5$
Copper II carbonate	$\text{CuCO}_3$
Sodium sulfate	$\text{Na}_2\text{SO}_4$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Mercury sulfate	$\text{HgSO}_4$
Nitrous acid	$\text{HNO}_2$
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron II nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron III nitrate	$\text{Fe}(\text{NO}_3)_3$
Carbon Dioxide	$\text{CO}_2$
Sulfur dioxide	$\text{SO}_2$
Sulfur pentoxide	$\text{SO}_5$
Tetraphosphorus octasulfide	$\text{P}_4\text{S}_8$
Carbonic acid	$\text{H}_2\text{CO}_3$

- Name the following compound and identify its molecular ions or acid.
  - $\text{Ca}(\text{NO}_3)_2$ , calcium (II) nitrate - II
  - $\text{Fe}(\text{NO}_3)_3$ , Iron (III) nitrate - III
  - $\text{Fe}(\text{OH})_2$ , Iron (II) hydroxide - II
  - $\text{N}_2\text{O}_5$ , Nitrogen pentoxide - III

# Chapter 9 Chemical Names And Formulas Assessment Answers

**Peter Atkins, Loretta Jones, Leroy  
Laverman**

## Chapter 9 Chemical Names And Formulas Assessment Answers:

**Essentials of Chemistry** Dennis D. Staley, Antony C. Wilbraham, 1984  
**Chemical Principles by Edward I. Peters, 2d Ed** Peter P. Berlow, 1978  
**Student's Guide to Introduction to Chemistry** Martha Mackin, 2012-12-02  
**Study Guide to Accompany Basics for Chemistry** is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *Essential Concepts of Chemistry* Sharon J. Sherman, 1998-09  
**Bulletin on Narcotics**, 1952  
**Loose-Leaf Version for Chemical Principles** Peter Atkins, Loretta Jones, Leroy Laverman, 2016-01-07 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry's frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do It also offers an exceptional level of support to help students develop their mathematical and problem solving skills For the new edition Chemical Principles now takes a modular approach with coverage organized as a series of brief Topics within 13 major areas of focus including a refresher on the fundamentals of chemistry and an online only section on techniques  
**General, Organic and Biochemistry** Kate Brown, William Henry Brown, 1983  
**Ebook: Physical Science** Tillery, 2016-04-16  
*Ebook Physical Science Introduction to Chemistry, Study Guide* T. R. Dickson, 1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining

problem solving skills Many chapters in this edition has been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images The New Science and Invention in Pictures ,1923 *Merrill Chemistry* Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 Science and Invention in Pictures ,1924 **Electrical Experimenter** ,1923 **The Medical World** ,1904 **10 in One Study Package for CBSE Science Class 10 with 3 Sample Papers & 16 Chapter Tests ebook** Disha Experts,2017-09-01 These books contain Access Codes along with instructions to access the Online Material In case you face any difficulty write to us at ebooks support aiets co in 10 in ONE CBSE Study Package Science class 10 with 3 Sample Papers is another innovative initiative from Disha Publication This book provides the excellent approach to Master the subject The book has 10 key ingredients that will help you achieve success 1 Chapter Utility Score CUS 2 Exhaustive Theory with Concept Maps 3 Text Book exercises 4 VSA SA LA Questions 5 Past year questions including 2017 Solved papers 6 HOTS Value based Exemplar 7 Past NTSE Exemplar MCQ s 8 16 Chapter Tests ebooks 9 Important Formulas Terms Definitions 10 3 Sample Papers with detailed solutions *Understanding Chemistry* Fred M. Dewey,1994 Prentice Hall Chemistry ,2000 Study Guide/Selected Solutions Manual Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book A Compact & Comprehensive Book of IIT Foundation Phy. & Che Class 7 S. K. Gupta & Anubhuti Gangal, Contains large number of Solved Examples and Practice Questions Answers Hints and Solutions have been provided to boost up the morale and increase the confidence level Self Assessment Sheets have been given at the end of each chapter to help the students to assess and evaluate their understanding of the concepts **Radio News** ,1923 Some issues 1943 July 1948 include separately paged and numbered section called Radio electronic engineering edition called Radionics edition in 1943

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