

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_ a. ionic bond  
 \_\_\_\_\_ **bond energy** \_\_\_\_\_ b. covalent bond  
 \_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_ a. H—F \_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_ d. H—H  
 \_\_\_\_\_ **ionic** \_\_\_\_\_ b. Na—Cl \_\_\_\_\_ **polar covalent** \_\_\_\_\_ e. H—C  
 \_\_\_\_\_ **polar covalent** \_\_\_\_\_ c. H—O \_\_\_\_\_ **polar covalent** \_\_\_\_\_ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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**Katrin Zwirglmaier**



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