

## Answer Key

### Chapter 7

#### Practice Problems

- neutral Br:  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^5$ ; ion:  $\text{Br}^-$ , anion,  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$
- neutral Ga:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^1$ ; ion:  $\text{Ga}^{3+}$ , cation,  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10}$
- neutral S:  $1s^2 2s^2 2p^6 3s^2 3p^4$ ; ion:  $\text{S}^{2-}$ , anion,  $1s^2 2s^2 2p^6 3s^2 3p^6$
- neutral Rb:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ ; ion:  $\text{Rb}^+$ , cation,  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$
- Al:  $[\text{Ne}]3s^2 3p^1$ ; F:  $[\text{He}]2s^2 2p^5$ ; one Al for every three F
- Li:  $[\text{He}]2s^1$ ; O:  $[\text{He}]2s^2 2p^4$ ; two Li for every O
- Be:  $[\text{He}]2s^2$ ; Se:  $[\text{Ar}]4s^2 3d^{10} 4p^4$ ; one Be for every Se
- Ga:  $[\text{Ar}]4s^2 3d^{10} 4p^1$ ; S:  $[\text{Ne}]3s^2 3p^4$ ; two Ga for every three S
- a. not ionic      b. not ionic      c. ionic      d. ionic
- a. LiF      b. MgS      c. MgO
- $\text{Na}_2\text{S}$       19.  $\text{Ca}(\text{OH})_2$
- $\text{Mg}_3\text{N}_2$       20. ammonium iodide
- $\text{K}_3\text{P}$       21. sodium bromate
- $\text{BaF}_2$       22. magnesium nitrate
- $\text{AlN}$       23. potassium hydrogen sulfate
- $\text{Na}_3\text{PO}_4$       24. ammonium perchlorate
- $(\text{NH}_4)_2\text{CO}_3$       25. aluminum hypochlorite
- $\text{Al}_2(\text{CrO}_4)_3$       26. iron(II) fluoride
- not conclusively, because ionic solids also conduct electricity when melted
- malleability; no, because metals are malleable

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**IM Harris**



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