

Section Review

Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium
- Identify three stresses that can change the equilibrium position of a chemical system
- Explain what the value of K_{eq} indicates about the position of equilibrium

Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Châtelier's principle
- equilibrium constant (K_{eq})

Key Equation

$$K_{eq} = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

When $aA + bB \rightleftharpoons cC + dD$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

| | |
|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------|
| _____ is the ratio of the concentrations of products to reactants, each raised to the power of its coefficient in the balanced chemical equation. | 1. _____ |
| _____ is the principle that states that if a dynamic equilibrium is subjected to a change in concentration, temperature, or pressure, the equilibrium position will shift to counteract the change. | 2. _____ |
| _____ is the position of a chemical system at equilibrium. | 3. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products remain constant over time. | 4. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products are equal. | 5. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products are not equal. | 6. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products are changing. | 7. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products are constant. | 8. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products are not constant. | 9. _____ |
| _____ is the state of a chemical system in which the concentrations of reactants and products are not changing. | 10. _____ |

18 3 Reversible Reactions And Equilibrium Answer Key

Charles H. Corwin



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