

### Equilibrium Extra Practice Problems

1. Quantities of  $\text{PCl}_3$  and  $\text{Cl}_2$  were placed in a reaction chamber and heated to  $230^\circ\text{C}$  at 101.3 kPa. At equilibrium  $[\text{PCl}_3] = 0.235 \text{ M}$  and  $[\text{PCl}_5] = 0.174 \text{ M}$ . If the equilibrium constant for this reaction is 0.0205, then calculate the  $[\text{Cl}_2]$  at equilibrium. *(0.0277 M)*



2. 1.0 mole of  $\text{NH}_3$  was introduced into a 1.00 L reaction vessel at a certain temperature. At equilibrium 0.60 mols of  $\text{H}_2$  were found to be present. Calculate the equilibrium constant. *(0.12)*



3. Consider the hypothetical reaction:  $\text{A}(g) + 2\text{B}(g) \rightleftharpoons 2\text{C}(g)$

The initial concentrations of the reactants and products were as follows:  $[\text{A}] = 2.00 \text{ M}$ ;  $[\text{B}] = 3.00 \text{ M}$ ;  $[\text{C}] = 1.00 \text{ M}$ . At equilibrium, it was found that the  $[\text{C}]$  was 3.00 M. Calculate the equilibrium constant. *(9.00)*

4. Consider the hypothetical reaction:  $2\text{A}(g) + \text{B}(g) \rightleftharpoons \text{C}(g) + 3\text{D}(g)$

Given that the initial concentrations of A, B, C & D were 4.0 M, 3.5 M, 2.0 M and 5.5 M respectively, and that  $[\text{A}]$  at equilibrium was 6.0 M, calculate the equilibrium constant. *(0.096)*

5. Calculate the equilibrium concentrations of all substances involved in the reaction depicted below when 0.10 M  $\text{CO}(g)$  is mixed with 0.10 M  $\text{H}_2\text{O}(g)$  in a 1L reaction vessel. *(x = 0.067 M)*



6. The equilibrium constant for the reaction above at  $700^\circ\text{C}$  is 5.10. Calculate the equilibrium concentrations of all substances when 1.00 mole of each compound is initially placed in a 1.00 L flask. *(x = 0.396 M)*

7. The equilibrium constant for the reaction below is 4.0. If 2.00 moles of each substance is placed in a 1.00 L flask, calculate the concentrations of all substances once equilibrium is reached. *(x = 0.500 M)*



# 17d Reaction Rates And Equilibrium Extra Practice Problems

**John T. Donoghue**



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