

10-2 Practice Problems

1. What is the volume of the oxygen produced when 10.0 g of potassium chlorate ($KClO_3$) is decomposed by heating under the proper conditions?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

2. Potassium permanganate decomposes under heat to produce K_2O , MnO_2 , and O_2 . How many grams of potassium permanganate must be heated to produce 1.00 g of oxygen?

$$\frac{1.00 \text{ g } O_2}{32.0 \text{ g/mol}} \times \frac{2 \text{ mol } KMnO_4}{3 \text{ mol } O_2} \times 158.0 \text{ g/mol} = 3.27 \text{ g } KMnO_4$$

3. What is the volume of the oxygen produced when 10.0 g of potassium chlorate ($KClO_3$) is decomposed by heating under the proper conditions?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

4. In practice, 10.0 g of $KClO_3$ is heated in a closed system and 2.00 g of oxygen is produced. How many grams of potassium chlorate must be heated?

$$\frac{2.00 \text{ g } O_2}{32.0 \text{ g/mol}} \times \frac{2 \text{ mol } KMnO_4}{3 \text{ mol } O_2} \times 158.0 \text{ g/mol} = 6.55 \text{ g } KMnO_4$$

5. 10.0 g of potassium chlorate ($KClO_3$) are decomposed by heating under the proper conditions. How many grams of oxygen are produced?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 32.0 \text{ g/mol} = 6.55 \text{ g } O_2$$

6. In order to produce 10.0 g of oxygen by heating potassium chlorate ($KClO_3$), how many grams of potassium chlorate must be heated?

$$\frac{10.0 \text{ g } O_2}{32.0 \text{ g/mol}} \times \frac{2 \text{ mol } KMnO_4}{3 \text{ mol } O_2} \times 158.0 \text{ g/mol} = 6.55 \text{ g } KMnO_4$$

7. Answer

7. What is the volume of the oxygen produced when 10.0 g of potassium chlorate ($KClO_3$) is decomposed by heating under the proper conditions?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

8. In practice, 10.0 g of a certain substance is decomposed by heating under the proper conditions and 2.00 g of oxygen is produced. How many grams of potassium chlorate must be heated?

$$\frac{2.00 \text{ g } O_2}{32.0 \text{ g/mol}} \times \frac{2 \text{ mol } KMnO_4}{3 \text{ mol } O_2} \times 158.0 \text{ g/mol} = 6.55 \text{ g } KMnO_4$$

9. What is the volume of a certain substance, 10.0 g of which is decomposed by heating under the proper conditions?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

10. What is the volume of a certain substance, 10.0 g of which is decomposed by heating under the proper conditions?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

11. In 10.0 g of potassium chlorate is decomposed by heating under the proper conditions.

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 32.0 \text{ g/mol} = 6.55 \text{ g } O_2$$

12. What is the volume of a certain substance, 10.0 g of which is decomposed by heating under the proper conditions?

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

$$\frac{10.0 \text{ g } KClO_3}{149.0 \text{ g/mol}} \times \frac{3 \text{ mol } O_2}{2 \text{ mol } KClO_3} \times 22.7 \text{ L/mol} = 2.70 \text{ L } O_2$$

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Siegfried Engelmann



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