

ACS Physical Chemistry Thermochemistry Study Guide

12.	What does -w mean?	Work is being produced - spontaneous
13.	What does a positive sign for w mean?	Work is required - non-spontaneous
14.	Can a change be spontaneous in both directions?	No
15.	What is an extensive property?	Value depends on the amount of substance
16.	What does $\Delta G > 0$ mean?	Non-spontaneous process - requires work
17.	What does $\Delta G < 0$ mean?	Spontaneous process - produces work
18.	What does standard free energy of formation mean? ΔG_o	The free energy change that occurs when 1 mole of a compound is made from its elements
19.	What is the standard free energy of formation for an element in its standard state? ΔG_f	0
20.	What does reversing a reaction do to the standard free energy of formation? ΔG_o	Changes its sign
21.	Why are most exothermic reactions spontaneous?	Because the large negative ΔH /large energy release makes the free energy change negative.
22.	What happens when ΔH and ΔS have opposite signs?	The reaction occurs spontaneously at all temperatures or at none.

Acs Physical Chemistry Exam Study Guide

PT Brinkman



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