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## ACS ANALYTICAL CHEMISTRY STUDY GUIDE

- $k_f$ :  $k_f = k_f \cdot a$
  - EDTA titrations: 1. before eq EDTA binds to metal 1:1
2. At eq we assume that all metal ions can bind to EDTA and then use  $k_f$
- Potentiometry: an analytical method in which an electric potential difference (avoltage) of a cell is measured
  - Reducing agent: Donates electrons and becomes oxidized.
  - oxidizing agent: Accepts electrons and becomes reduced.
  - coulometry: science of measuring change and relating it to analyte
  - cathode: the electrode at which reduction occurs
  - Anode: the electrode at which oxidation occurs
  - $E_{cell}$ :  $E_{cell} = E_{red} - E_{ox}$
  - Faraday's Equation:  $E = E - (0.0592/n) \log(a/b)$
  - ||: anode always written on left of ||
  - amperometry: elec techniques where the cell current is measured where pot diff between the indicator and reference electrode is controlled
  - Clark electrode: Electrode for oxygen
  - Biosensors: Detecting devices that can detect specific substances in cells
  - amp chromo det: point, size, or charge
  - limiting current: the diff between current and steady state max
  - Anodic vol: Very sensitive for analyzing metals into ppb
  - fluorescence: used to measure singlets, atoms like rigid aromatic rings, high sensitivity
  - phosphorescence: involves emission of light trip- singlet, much lower than fluoro
  - Transmittance:  $A = -\log(T)$  or  $2 - \log(T\%)$
  - atomic absorption: lamp-flame-mono-detector
  - atomic emission: flame-mono-detector
  - AA hollow cathode lamp: used to form very narrow band lights
  - AA atomizer: Heat sample to very high temps to break all chemical bonds, hot enough to

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**H Kauffman**



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