

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H ⁺) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a <u>proton donor</u>	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H ⁺) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids	These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid Sulphuric acid Nitric acid	HCl H ₂ SO ₄ HNO ₃ MINERAL ACIDS	They are derived from mineral salts i.e chlorides for HCl, sulphates for H ₂ SO ₄ and nitrates for HNO ₃

Other mineral acids known	Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid Carbonic acid Phosphoric acid Nitrous acid	H ₂ SO ₃ H ₂ CO ₃ H ₃ PO ₄ HNO ₂ Derived from sulphites Derived from carbonates Derived from phosphates Derived from nitrites	Ethanoic acid (CH ₃ COOH) Methanoic acid (HCOOH)	CITRIC ACID from lemons TARTARIC ACID from grapes ACETIC ACID from vinegar LACTIC ACID from sour milk Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule of an acid in aqueous solution</u> .
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Acids Bases And Salts Study Guide

John McMurry, Susan McMurry



Acids Bases And Salts Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1992

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