

19.1 Acid-Base Theories [587]: Section Assessment

- Due Apr 23 at 11:59pm
- Points 16
- Questions 16
- Available Apr 18 at 12am - May 24 at 11:59pm
- Time Limit None
- Allowed Attempts 3

Take the Quiz Again

Attempt History

	Attempt	Time	Score
LATEST	Attempt 1	1,426 minutes	14 out of 16

① Answers will be shown after your last attempt

Score for this attempt: 14 out of 16

Submitted Apr 19 at 4:19pm

This attempt took 1,426 minutes.



Question 1

1 / 1 pts

Write a chemical equation for the ionization of HNO_3 in water.

- A) $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{NO}_3^-$
- B) $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O} + \text{NO}_3^-$
- C) $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O} + \text{NO}_3$
- D) $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{OH}^- + \text{H}_2\text{NO}_3^+$

- D
- C
- B
- A



Question 2

1 / 1 pts

Write a chemical equation for the reaction of CO_3^{2-} with water.

Acid Base Theories Section Review Answer

Raymond E. Davis, Kenneth W. Whitten

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