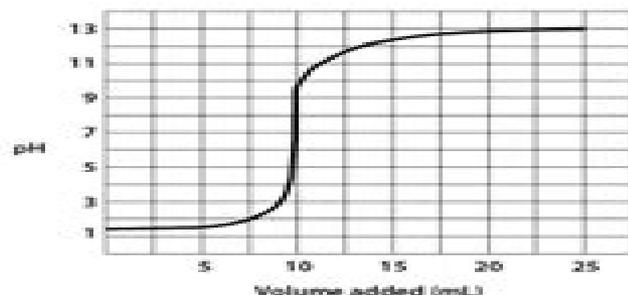
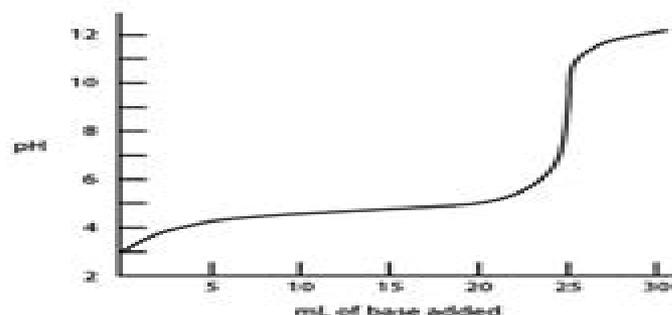


# Acid-Base Titration Curve

(1) Consider this titration curve.

- This is a \_\_\_\_\_ (strong/weak acid) titrated with a strong base. The acid is \_\_\_\_\_ (monoprotic/diprotic).
- Place a dot (•) at the equivalence point. The pH at the equivalence point is \_\_\_\_\_.
- What volume of the base was used to titrate the acid solution? \_\_\_\_\_ mL.
- Place a box (■) on the curve where the pH of the solution =  $pK_a$  of the acid.
  - What is the pH at this point? \_\_\_\_\_.
  - What is the  $pK_a$  of the acid? \_\_\_\_\_.
  - What is the  $K_a$  of the acid? \_\_\_\_\_.



(2) From the given titration curve, determine the following:

- if the acids/bases are strong or weak
  - the volume required to reach the equivalence point
- whether the acid or base was added from the buret
  - the pH at the equivalence point
  - the concentration of base used if this titration started with 50.0 mL of 0.050M acid.

# Acid Base Titration Worksheet With Answers

**Sally Diane Rupert**



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**First Year Chemistry Students' Conceptions of Acid/base Chemistry** Sally Diane Rupert, 2001      **Experimental Chemistry** Robert J. Artz, 1982      **Educart ICSE Class 10 One-shot Question Bank 2026 Chemistry (strictly for 2025-26 boards)** Sir Tarun Rupani, 2025-07-12

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**Acid-base Titrations in Nonaqueous Solvents** James Sherwood Fritz, 1952      **Aqueous Acid-base Equilibria and Titrations** Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses new theoretical developments which have led to more

generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved

**Acid-Base Diagrams** Heike Kahlert, Fritz Scholz, 2013-07-31 Understanding acid base equilibria made easy for students in chemistry biochemistry biology environmental and earth sciences Solving chemical problems be it in education or in real life often requires the understanding of the acid base equilibria behind them Based on many years of teaching experience Heike Kahlert and Fritz Scholz present a powerful tool to meet such challenges They provide a simple guide to the fundamentals and applications of acid base diagrams avoiding complex mathematics This textbook is richly illustrated and has full color throughout It offers learning features such as boxed results and a collection of formulae

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**Acid-bases in Analytical Chemistry** Izaak Maurits Kolthoff,1966      **General Theory of Acid-base Titration** Pat Richard Pondy,1949      **Volumetric Analysis** Ralf Martens-Menzel,Lena Harwardt,Hanns-Jürgen Krauss,2025-05-19 This standard work on volumetric analysis based on the 20th German edition provides comprehensive information on the theory of acid base titration redox titration complexation titration and precipitation titration with both classical and instrumental indication of the equivalence point Many applications are described and explained in detail with examples in pharmaceutical and environmental analysis

*Working with Chem Separate* Wink,1999-12-01      **Solving Acid-Base Titration and Equilibrium Problems Using Computer Algebra Software** Andrew Biaglow,2008-12-01 This book is a treatment of several simple chemical equilibrium problems using charge and mole balances and equilibrium expressions We show how to arrive at an algebraic solution without the assumptions common in general chemistry textbooks The methods involve application of Mathematica or Maple software to solve the problems with minimal algebraic work on the part of the student

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