

## AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

### Unit 8 – Introduction to Equilibrium

# KEY



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



- (A) 95<sup>2</sup>  
 (B) 95  
 (C) (95)<sup>2</sup>  
 (D)  $\frac{1}{95}$   
 (E)  $\frac{1}{(95)^2}$



$C_2H_2(g)$  is reduced by  $H_2(g)$  in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of  $H_2(g)$   
 II. Changes in the particle size of the platinum catalyst  
 III. Changes in the temperature of the reaction system

- (A) III only  
 (B) I and II only  
 (C) I and III only  
 (D) II and III only  
 (E) I, II, and III



A sealed rigid vessel contains  $BaO_2(s)$  in equilibrium with  $BaO(s)$  and  $O_2(g)$  as represented by the equation above. Which of the following changes will increase the amount of  $BaO_2(s)$  in the vessel?

- (A) Removing a small amount of  $O_2(g)$  *R*  
 (B) Removing a small amount of  $BaO(s)$  *X*  
 (C) Adding He gas to the vessel *X*  
 (D) Lowering the temperature *L*

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? *increase P*

- (A)  $SiCO_3(l) \rightleftharpoons SiO_3(l) + CO_2(g)$   
 (B)  $Cl_2 + COCl_2(g) \rightleftharpoons 2 COCl(g)$   
 (C)  $CO(g) + H_2O(g) \rightleftharpoons CO_2(g) + H_2(g)$   
 (D)  $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$   
 (E)  $2 NO_2(g) \rightleftharpoons N_2O_4(g)$
- shift towards fewer & make gas*



When 4.00 mol each of  $X(g)$  and  $Y(g)$  are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of  $Z(g)$  is produced. What is the value of the equilibrium constant,  $K_c$ ?

- (A) 3  
 (B) 6  
 (C) 8  
 (D) 16  
 (E) 36
- $X + Y \rightleftharpoons 2Z$   
 $4M \quad 4M \quad 0$   
 $-x \quad -x \quad +2x$   
 $1M \quad 1M \quad 6M$   
 $2x = 6M \Rightarrow x = 3M$   
 $K_c = \frac{[Z]^2}{[X][Y]} = \frac{6^2}{1 \times 1} = 36$

6. The synthesis of  $CH_3OH(l)$  from  $CO(g)$  and  $H_2(g)$  is represented by the equation above. The value of  $K_c$  for the reaction at 483 K is 14.5.

Which of the following explains the effect on the equilibrium constant,  $K_c$ , when the temperature of the reaction system is increased to 650 K?

- (A)  $K_c$  will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.  
 (B)  $K_c$  will increase because there are more reactant molecules than product molecules.  
 (C)  $K_c$  will decrease because the reaction is exothermic.  
 (D)  $K_c$  is constant and will not change.

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