

# Science

(www.tiwariacademy.com)  
(Chapter – 1) (Matter in Our Surroundings)  
(Class – IX)

## Question 4:

Liquids generally have lower density as compared to solids. But you must have observed that ice floats on water. Find out why.

### Answer 4:

The mass per unit volume of a substance is called density (density = mass/volume).

*As the volume of a substance increases, its density decreases.*

Though ice is a solid, it has large number of empty spaces between its particles. These spaces are larger as compared to the spaces present between the particles of water. Thus, the volume of ice is greater than that of water. Hence, the density of ice is less than that of water. A substance with lower density than water can float on water. Therefore, ice floats on water.

## Page 9

## Question 1:

Convert the following temperatures to the celsius scale.

(a) 300 K

(b) 573 K

### Answer 1:

(a)  $300\text{ K} = (300 - 273)^\circ\text{C} = 27^\circ\text{C}$

(b)  $573\text{ K} = (573 - 273)^\circ\text{C} = 300^\circ\text{C}$

## Question 2:

Convert the following temperatures to the celsius scale.

(a)  $250^\circ$

(b)  $100^\circ$

### Answer 2:

(a) Water at  $250^\circ\text{C}$  exists in gaseous state.

(b) At  $100^\circ\text{C}$ , water can exist in both liquid and gaseous form. At this temperature, after getting the heat equal to the latent heat of vaporization, water starts changing from liquid state to gaseous state.

## Question 3:

For any substance, why does the temperature remain constant during the change of state?

### Answer 3:

During a change of state, the temperature remains constant. This is because all the heat supplied to increase the temperature is utilised (as latent heat) in changing the state by overcoming the forces of attraction between the particles. Therefore, this heat does not contribute in increasing the temperature of the substance.

## Question 4:

Suggest a method to liquefy atmospheric gases.

### Answer 4:

By applying pressure and reducing the temperature, atmospheric gases can be liquefied.

## Page 10

## Question 1:

Why does a desert cooler cool better on a hot dry day?

### Answer 1:

When a liquid evaporates, the particles of the liquid absorb energy from the surroundings to compensate the loss of energy during evaporation. This makes the surroundings cool.

# 2015wacce Science Question

**F Rizvi**



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## **Table of Contents 2015wacce Science Question**

1. Understanding the eBook 2015wacce Science Question
  - The Rise of Digital Reading 2015wacce Science Question
  - Advantages of eBooks Over Traditional Books
2. Identifying 2015wacce Science Question
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2015wacce Science Question
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2015wacce Science Question
  - Personalized Recommendations
  - 2015wacce Science Question User Reviews and Ratings
  - 2015wacce Science Question and Bestseller Lists
5. Accessing 2015wacce Science Question Free and Paid eBooks
  - 2015wacce Science Question Public Domain eBooks
  - 2015wacce Science Question eBook Subscription Services
  - 2015wacce Science Question Budget-Friendly Options
6. Navigating 2015wacce Science Question eBook Formats

- ePub, PDF, MOBI, and More
  - 2015wacce Science Question Compatibility with Devices
  - 2015wacce Science Question Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 2015wacce Science Question
  - Highlighting and Note-Taking 2015wacce Science Question
  - Interactive Elements 2015wacce Science Question
8. Staying Engaged with 2015wacce Science Question
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2015wacce Science Question
9. Balancing eBooks and Physical Books 2015wacce Science Question
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2015wacce Science Question
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 2015wacce Science Question
- Setting Reading Goals 2015wacce Science Question
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2015wacce Science Question
- Fact-Checking eBook Content of 2015wacce Science Question
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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