

2023 MCQs IEP



Based on the chemical equations given, which of the following statements is correct?

- (A) **HCN** is a weaker acid than **H₂O**.
 (B) **CN⁻** is a stronger acid than **HCN**.
 (C) **CN⁻** is a stronger base than **H₂O**.
 (D) **H₂O** is a weaker base than **HCN**.



A 0.20 mol sample of **H₂S(g)** and a 0.10 mol sample of **SO₂(g)** are placed in a previously evacuated 10.0 L vessel at 298 K. Even though the reaction represented by the equation above is thermodynamically favorable, the formation of **S(s)** was not observed. Which of the following best explains why **S(s)** is not produced at 298 K?

- (A) At 298 K, some **S(s)** must first be present for the reaction to occur.
 (B) At 298 K, the value of the equilibrium constant is less than 1.
 (C) At 298 K, the energy of most molecular collisions is lower than the activation energy for the reaction.
 (D) At 298 K, the loss of entropy is too great to make molecular collisions likely to occur.

3.

Half-Reaction	E° (V)
$\text{Ag}^+(aq) + e^- \rightarrow \text{Ag}(s)$	0.80
$\text{Ni}^{2+}(aq) + 2 e^- \rightarrow \text{Ni}(s)$	-0.25

The standard reduction potentials of the half-reactions associated with a galvanic cell are listed in the table above. Which of the following statements identifies the electrode that increases in mass as the cell operates, and correctly explains why?

- (A) The **Ag** electrode, because **Ag⁺(aq)** in the solution is reduced to form **Ag(s)**.
 (B) The **Ag** electrode, because **Ag** has a greater molar mass than **Ni** has.
 (C) The **Ni** electrode, because **Ni** is chemically more reactive than **Ag** is.
 (D) The **Ni** electrode, because **Ni²⁺(aq)** is reduced at a faster rate than **Ag⁺(aq)** is.

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