

ChemThink: Covalent Bonding

Consider bonding between two atoms whose outer shells electrons

When two atoms get closer together, the electrons attract the other atom's

Protons & Neutrons & Electrons

Before bonding, the atoms's electrons spend most of their time

near the nuclei of each atom. Once bonded, the electrons spend most of their time between the two nuclei

atoms must be able to hold onto their own electrons while giving another atom's electrons.

Consider bonding between two electrons

When atoms move closer, the potential energy increases & decreases

At a certain point the potential energy decreases if you keep
and lower the atomic number than protons to which
nucleus are repelling each other.

This ideal distance between the atoms is known as the bond length

Lower in energy = more stable

Bond Type	Example	# of shared e ⁻	Total # of e ⁻ shared	Strength
Single	H—H	1	2	weakest
Double	O=O	2	4	stronger
Triple	N≡N	3	6	strongest

Chemthink Covalent Bonds

VM Jensen

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular glow behind it.

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