

## Introduction 100

# pH and pOH

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### Introduction Questions

- What are the units on pH and pOH? What are general practices relating the logarithm base 10's coefficient? Differentiate  $\log_{10}$  concentration from pH.
  - Take the log with three significant figures. How about pH = 0.1234?
  - How many more than the logarithm value does your pH have? (e.g., pH = 0.1234)
- What are the units on  $K_w$ ? (e.g.,  $10^{-14}$ )
- Calculate the following:
 

|                     |                    |                 |
|---------------------|--------------------|-----------------|
| a. $\log(10^{-14})$ | b. $\log(10^{-1})$ | c. $\log(10^0)$ |
| 14                  | -1                 | 0               |

### Introduction: Calculating pH

Many properties based on  $[H^+]$  and many people know that pH is a logarithm. But what does mean? Calculated what activity is measured. For pH the standard concentration of hydrogen ions  $[H^+]$  is 1 mol/L.

What pH is solution with concentration of  $1.0 \times 10^{-7}$  mol/L calculating pH is:

$$pH = -\log(10^{-7})$$

Therefore, the pH of a solution with the concentration of  $1.0 \times 10^{-7}$  mol/L is 7. The calculated pH is 7.00. The pH is 7.00 and there is no error.

### Critical Thinking Questions

- In a concentration base concentration of hydrogen ions of  $1.0 \times 10^{-7}$  mol/L, what is the pH?
  - Calculating  $-\log(10^{-7}) = 7.00$
- In a concentration base concentration of  $1.0 \times 10^{-7}$  mol/L, what is the pH?
  - What is the concentration of  $[H^+]$  ions?
    - Calculating  $10^{-7}$

# Chemquest Ph And Poh Answer Key

**Michal Rosen-Zvi**



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