

Electron Configurations and the Periodic Table

Name: _____

Date: _____

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Information: Relating Electron Configurations to the Periodic Table

In this section you will see how the periodic table serves as a road map for writing electron configurations. Get your periodic table out and get ready. Remember that a *row* on the periodic table goes horizontally from left to right. *Columns* are vertical (up and down).

Critical Thinking Questions

- Write the electron configurations for Li, Na and K. (Remember for electron configurations, atoms are not necessary.)
- What is similar about the electron configurations of all of the elements in question 1? Look at the way end of their configurations.
- Lithium (Li) is in row 2 of the periodic table, sodium (Na) is in row 3, and potassium is in row 4. How do their row numbers affect how their electron configurations end?
- Write the electron configurations for Be, Mg, and Ca.
- What is similar about the ending of the electron configurations for all of the elements in question 4?
- Beryllium (Be) is in row 2 of the periodic table, magnesium (Mg) is in row 3, and Calcium (Ca) is in row 4. How do their row numbers affect how their electron configurations end?
- Given what you have done in questions 1-6, complete the following statement:
Elements ending in s^1 are in column number _____ and those ending in s^2 are in column number _____ of the periodic table.

- Name the element that have an electron configuration ending with... (the first is done for you)
a) $3s^1$ Lithium (Li) b) $4s^2$ _____ c) $3s^2$ _____
- Write the electron configurations for B, Al, and Ga and note their similarities.
- Write the electron configurations for N, P and As and note their similarities.
- Complete the following statement: Elements ending in p^1 are in column number _____ of the periodic table and elements ending in p^5 are in column number _____ of the periodic table. Therefore, elements ending in p^3 must be in column _____ of the periodic table.
- Name the elements that have an electron configuration ending with... (the first is done for you)
a) $3p^2$ Sulfur (S) b) $3p^4$ _____ c) $3p^1$ _____
- Write the electron configurations for Ti, Zr, and Hf and note their similarities.
- Write the electron configurations for Cr, Mo, and W and note what is similar about them.
- Complete the following: Elements ending in d^1 are in column number _____ and those ending in d^9 are in column number _____. Therefore, elements ending in d^5 are in column number _____. Elements ending in d^7 must be in column number _____.
- Notice from question 8 that an element that ends in $3s$ is in the 3rd row. Also, from questions 9-12, it should be clear that an element that ends in $3p$ is in the 3rd row. However, notice that an element that ends in $3d$ is in the 4th row instead of the 3rd row. Offer an explanation for this.
- Name the elements that have an electron configuration ending with... (the first is done for you)
a) $4d^2$ Niobium (Nb) b) $5d^2$ _____ c) $4d^1$ _____

Chemquest 20 Worksheet

Albert A Gayle



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