

IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1. $\text{K} + \text{F}$



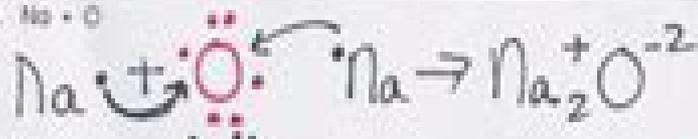
2. $\text{Mg} + \text{I}$



3. $\text{Be} + \text{S}$



4. $\text{Na} + \text{O}$



5. $\text{Al} + \text{Br}$



COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:

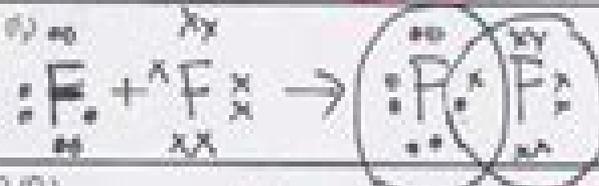


Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1. $\text{H} + \text{H} (\text{H}_2)$



2. $\text{F} + \text{F} (\text{F}_2)$



3. $\text{O} + \text{O} (\text{O}_2)$



4. $\text{N} + \text{N} (\text{N}_2)$



5. $\text{C} + \text{O} (\text{CO}_2)$



6. $\text{H} + \text{O} (\text{H}_2\text{O})$



Chemquest 20 Advanced Ionic Bonding Answer Key

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