

Chemistry – Unit 5

Workshop 5 – Gases Again

1. A metal powder reacts completely with oxygen gas at 100 °C. The products of the gas are 1.170 g of M_2O when the temperature is 100 °C. The total volume of the gas (metal + oxygen) at 100 °C is 1.00 dm³ at the pressure of the gas. What is the relative atomic mass of the metal?
2. A 10.0 g sample of sodium metal reacts with oxygen gas at 100 °C. An 11.0 g solid product is formed. The products of the reaction are 0.100 dm³ of Na_2O and 0.100 dm³ of Na_2O_2 . Assume oxygen gas was collected at 100 °C and 1.00 atm. Calculate the relative atomic mass of the metal. Is the metal sodium?
3. A small bubble containing methane is released from the bottom of a 10.0 m deep well. When the temperature is 20 °C and the pressure is 1.00 atm, the bubble's volume is 1.00 cm³. Assuming the temperature is 20 °C and the pressure is 1.00 atm, calculate the mass of methane gas in the bubble. Is the bubble of the same volume as the one at the top?
4. Three gases are mixed in a 1.00 dm³ container. The partial pressures of O_2 , N_2 and CO_2 are 1.00 atm, 1.00 atm and 1.00 atm. What is the pressure of the mixture of gases?
5. What are the percentages by volume of the gases in the above mixture?

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