

Chemistry – Unit 8 Worksheet 4

1. How many kJ of heat will be released when 4.72 g of carbon react with excess oxygen gas to produce carbon dioxide?



$$\Delta H^\circ = -393.5 \text{ kJ}$$

2. How much heat should be transferred when 38.2 g of liquid bromine reacts with excess hydrogen gas to form hydrogen bromide? Is the heat being transferred from the system to the surroundings or from the surroundings to the system?



$$\Delta H^\circ = -72.80 \text{ kJ}$$

3. How many kJ of heat would you expect to be transferred when 6.44 g of sulfur react with excess oxygen gas to produce sulfur trioxide? Is this reaction endothermic or exothermic?



$$\Delta H^\circ = -791.4 \text{ kJ}$$

4. Nitrogen gas and oxygen gas can combine to produce nitric oxide, NO. If such a reaction absorbs 88.0 kJ of heat from the surroundings, how many grams of nitrogen gas do you predict were consumed in the reaction?



$$\Delta H^\circ = 180 \text{ kJ}$$

Chemistry Unit 8 Ws 2 Answers

Charles H. Corwin



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