

Answer Key - AP Chemistry Summer Assignment # 1

1. How many significant figures are there in each of the following values?

| | | | |
|--------------|-----------------------|---------------|-----------------------|
| a.) 0.002330 | 4 significant figures | e.) 0.0002 | 1 significant figure |
| b.) 13.00 | 4 significant figures | f.) 2200.0 | 5 significant figures |
| c.) 522.1221 | 7 significant figures | g.) 0.0531120 | 6 significant figures |
| d.) 1204.30 | 6 significant figures | | |

2. Perform the indicated calculations on the following measured values, giving the final answer with the correct number of significant figures.

| | |
|---|---|
| a.) $16.81 + 3.2257 = 20.0357 \approx 20.04$ | b.) $324.6 \times 815.991 = 264870.6786 \approx 264900$ |
| c.) $2.85 + 3.4621 + 1.3 = 7.6121 \approx 7.6$ | d.) $7.442 - 7.429 = 0.013$ |
| e.) $1.65 \times 14 = 23.1 \approx 23$ | f.) $27 / 4.148 = 6.509161... \approx 6.5$ |
| g.) $[(3.901 - 3.887) / 3.901] \times 1.00 = [0.014 / 3.901] \times 1.00 = 0.0036 \times 1.00 = 0.0036$ | |
| h.) $6.404 \times 2.91 \times (18.7 - 17.1) = 6.404 \times 2.91 \times 1.6 \approx 30.$ | |

3. A sample of motor oil with a mass of 440 g occupies 500 mL. What is the density of the motor oil?

$$d = \frac{m}{v} \rightarrow d = \frac{440 \text{ g}}{500 \text{ mL}} \rightarrow d = 0.88 \text{ g/mL} \rightarrow d = 0.9 \text{ g/mL}$$

4. The density of an object is 16.3 g/mL. Its volume is 0.125 L. What is the mass of the object?

$$\frac{16.3 \text{ g}}{1 \text{ mL}} \left| \frac{1000 \text{ mL}}{1 \text{ L}} \right. = 16300 \text{ g/L}$$

$$d = \frac{m}{v} \rightarrow 16300 \text{ g/L} = \frac{m}{0.125 \text{ mL}} \rightarrow m = 16300 \times 0.125 \rightarrow m = 2037.5 \text{ g} \rightarrow m = 2040 \text{ g}$$

5. A sample of uranium weighing 30.923 g was dropped in a graduated cylinder containing 22.30 mL of water. The volume of the water plus the sample was 23.90 mL. What is the density of uranium?

$$\text{Volume} = 23.90 \text{ mL} - 22.30 \text{ mL} = 1.60 \text{ mL}$$

$$d = \frac{m}{v} \rightarrow d = \frac{30.923 \text{ g}}{1.60 \text{ mL}} \rightarrow d = 19.33 \text{ g/mL} \rightarrow d = 19.3 \text{ g/mL}$$

6. How many protons, neutrons and electrons are in each of the following ions?

| | |
|---|---|
| a.) mass #56, atomic #26, 3+ charge 26 protons $56 - 26 = 30$ neutrons $26 - 3 = 23$ electrons | b.) mass #40, atomic #20, 2+ charge 20 protons $40 - 20 = 20$ neutrons $20 - 2 = 18$ electrons |
| c.) mass #19, atomic #9, 1- charge 9 protons $19 - 9 = 10$ neutrons $9 + 1 = 10$ electrons | d.) mass #31, atomic #15, 3- charge 15 protons $31 - 15 = 16$ neutrons $15 + 3 = 18$ electrons |
| e.) mass #127, atomic #53, 1- charge 53 protons $127 - 53 = 74$ neutrons $53 + 1 = 54$ electrons | f.) mass #127, atomic #53, 7+ charge 53 protons $127 - 53 = 74$ neutrons $53 - 7 = 46$ electrons |

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Raffaella Di Napoli



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