

Chapter 6

Chapter 6: Thermochemistry

6.2 - Enthalpy and Calorimetry

Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{\text{products}} - H_{\text{reactants}}$$

Products of a reaction have a greater enthalpy (ΔH):

- ΔH is positive
- heat is absorbed by the system (endothermic reaction)

Reactants of a reaction have a greater enthalpy (ΔH):

- ΔH is negative
- heat is released by the system (exothermic reaction)

Calorimetry

Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

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McGraw-Hill Staff, 2001-03 **Organic Chemistry, International Adaptation** T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2023-01-19 Organic Chemistry 13th edition provides a comprehensive yet accessible treatment of all the essential organic chemistry concepts with emphasis on relationship between structure and reactivity in the subject The textbook includes all the concepts covered in a typical organic chemistry textbook but is unique in its skill development approach to the subject Numerous hands on activities and real world examples are integrated throughout the text to help students understand both the why and the how behind organic chemistry This International Adaptation offers new and updated content with improved presentation of all course material It offers new material on several topics including the relevance of intermolecular forces in the immune response and vaccines like those for Covid 19 the chemistry of breathing carbonic anhydrase how conjugation and complexation affect the color of lobsters and how biodegradable polymers are used to stabilize vaccines and pharmaceuticals Content is revised to reflect the current understanding of chemical processes and improved depictions of longstanding mechanisms This edition builds on the ongoing pedagogical strength of the book with the inclusion of additional worked and end of chapter problems and an engaging set of new problems entitled Chemical Consultant Needed These draw from the primary chemical literature and give students experience of working with more complex polyfunctional structures and areas where key transformations take place **Organic Chemistry** Francis A.

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those teaching chemistry across educational levels including both those intended to replace teaching from the front and others that can be built into traditional lecture courses to enhance the learning experience

Understanding Organic Chemistry Margot K. Schumm, 1987

The Success Manual for General Chemistry Elizabeth Kean, Catherine Middlecamp, 1986

Student-involved Assessment for Learning Richard J. Stiggins, 2005 This text for assessment courses focuses on showing prospective teachers how to develop assessments that accurately reflect student achievement and how to use those assessments to benefit not merely grade student learning A variety of hands on practice activities provide clear guidance on how to construct all types of assessment items and tests

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Resources in Education, 1994-02

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