



Define the following terms.

1. atomic number (Z): The number of protons in the nucleus.
2. mass number: The number of protons and neutrons in the nucleus.
3. isotope: Atoms of the same element (same number of protons) having differing numbers of neutrons.
4. nuclide: Atoms of a given element which have the same number of neutrons.

Answer the following questions.

5. What is the result if the atomic number of a nuclide is subtracted from the mass number of that nuclide?

The number of neutrons.

Look at Table 4.3 on the top of page 114. Write the symbol for each of the following:

6. proton p or p^+
7. neutron n or n^0
8. electron e^-

In the third paragraph of page 117, find the name of the method used to describe an isotope when written as copper-63:

9. isotope notation (or hyphen notation)

Write the nuclear symbol method for showing copper-63 that shows the chemical symbol, mass number, and atomic number as shown in figure 4-16 at the top of page 117.

10. ${}^{63}_{29}\text{Cu}$

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