

7. molar heat of condensation

### Part B True-False

8. AT                      10. ST                      12. NT  
9. NT                      11. NT

### Part C Matching

13. a                      15. b                      17. d  
14. e                      16. c

### Part D Questions and Problems

18. a. endothermic                      d. endothermic  
b. endothermic                      e. exothermic  
c. exothermic                      f. exothermic

$$19. \Delta H = 28.3 \frac{\text{g H}_2\text{O}(s)}{1 \text{ mol H}_2\text{O}(s)} \times \frac{1 \text{ mol H}_2\text{O}(s)}{18.0 \text{ g H}_2\text{O}(s)} \\ \times \frac{6.01 \text{ kJ}}{1 \text{ mol H}_2\text{O}(s)} = 9.45 \text{ kJ}$$

$$20. \Delta H = 5.53 \frac{\text{mol NH}_2\text{NO}_2(s)}{1 \text{ mol NH}_2\text{NO}_2(s)} \\ \times \frac{25.7 \text{ kJ}}{1 \text{ mol NH}_2\text{NO}_2(s)} = 142 \text{ kJ}$$

## Section 17.4

### Part A Completion

- sum
- enthalpy
- indirectly
- changed (reversed)
- standard heat of formation
- change
- one
- $\Delta H_f^\circ$
- zero
- subtracting

### Part B True-False

11. NT                      13. AT                      15. NT  
12. NT                      14. NT

### Part C Matching

16. b                      18. a                      20. c  
17. c                      19. d

### Part D Questions and Problems



$$22. \Delta H_f^\circ (\text{products}) \\ = 3 \frac{\text{mol CO}_2(g)}{1 \text{ mol CO}_2(g)} \times \frac{-393.5 \text{ kJ}}{1 \text{ mol CO}_2(g)} \\ = -1181 \text{ kJ}$$

$$\Delta H_f^\circ (\text{reactants}) \\ = 3 \frac{\text{mol CO}(g)}{1 \text{ mol CO}(g)} \times \frac{-110.5 \text{ kJ}}{1 \text{ mol CO}(g)} + (-822.1 \text{ kJ}) \\ = -1154 \text{ kJ}$$

$$\Delta H^\circ = \Delta H_f^\circ (\text{products}) - \Delta H_f^\circ (\text{reactants}) \\ = -1181 \text{ kJ} - (-1154 \text{ kJ}) \\ = -27 \text{ kJ}$$

## Practice Problems

### Section 17.1

- $200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}} \\ = 836.8 \text{ kJ}$
- $C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0^\circ\text{C}} = 1.40 \text{ cal}/(\text{g} \times ^\circ\text{C})$
- $\Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J}/(\text{g} \times ^\circ\text{C})} = 6.0^\circ\text{C}$
- $q = 100.0 \text{ g} \times 120.0^\circ\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \\ = 1.1 \times 10^4 \text{ J}$

### Section 17.2

- $\Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \times 10^\circ\text{C} \\ = 6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ}$
- $\Delta H = 15.0 \text{ g Ca(OH)}_2(s) \times \frac{1 \text{ mol Ca(OH)}_2(s)}{74.1 \text{ g Ca(OH)}_2(s)} \\ \times \frac{-65.2 \text{ kJ}}{1 \text{ mol Ca(OH)}_2(s)} = -13.2 \text{ kJ}$
- $\Delta H = 52.4 \text{ g CH}_4(g) \times \frac{1 \text{ mol CH}_4(g)}{16.0 \text{ g CH}_4(g)} \\ \times \frac{-890.2 \text{ kJ}}{1 \text{ mol CH}_4(g)} = 2.93 \times 10^3 \text{ kJ}$
- $4\text{NH}_3 + 5\text{O}_2(g) \rightarrow 4\text{NO}(g) + 6\text{H}_2\text{O}(g) \\ \Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_3(g)} \times 4 \text{ mol NH}_3(g) \\ = -904 \text{ kJ}$

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1. Understanding the eBook Chemistry Prentice Hall Study Guide Answer Key
  - The Rise of Digital Reading Chemistry Prentice Hall Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Prentice Hall Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Prentice Hall Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Prentice Hall Study Guide Answer Key
  - Personalized Recommendations
  - Chemistry Prentice Hall Study Guide Answer Key User Reviews and Ratings
  - Chemistry Prentice Hall Study Guide Answer Key and Bestseller Lists
5. Accessing Chemistry Prentice Hall Study Guide Answer Key Free and Paid eBooks
  - Chemistry Prentice Hall Study Guide Answer Key Public Domain eBooks
  - Chemistry Prentice Hall Study Guide Answer Key eBook Subscription Services
  - Chemistry Prentice Hall Study Guide Answer Key Budget-Friendly Options

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  - ePub, PDF, MOBI, and More
  - Chemistry Prentice Hall Study Guide Answer Key Compatibility with Devices
  - Chemistry Prentice Hall Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Prentice Hall Study Guide Answer Key
  - Highlighting and Note-Taking Chemistry Prentice Hall Study Guide Answer Key
  - Interactive Elements Chemistry Prentice Hall Study Guide Answer Key
8. Staying Engaged with Chemistry Prentice Hall Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Prentice Hall Study Guide Answer Key
9. Balancing eBooks and Physical Books Chemistry Prentice Hall Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Prentice Hall Study Guide Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Prentice Hall Study Guide Answer Key
  - Setting Reading Goals Chemistry Prentice Hall Study Guide Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Prentice Hall Study Guide Answer Key
  - Fact-Checking eBook Content of Chemistry Prentice Hall Study Guide Answer Key
  - Distinguishing Credible Sources
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  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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