

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

| | | |
|---------------|---------------------------|---------------------------------|
| 14.8 moles Al | 3 moles of O ₂ | |
| | 4 mole Al | = 11.1 mol O₂ |

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO₃).



| | | |
|--------------------------------|---------------------------|---|
| 0.61 moles KClO ₃ | 3 moles of O ₂ | |
| | 2 mole KClO ₃ | = |
| 0.915 mol O₂ | | |

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl₂) are produced?



| | | | |
|----------------------------------|----------|--------------------------|-------------------------|
| 20 g Zn | 1 mol Zn | 1 mole ZnCl ₂ | 136 g ZnCl ₂ |
| | 65g Zn | 1 mol Zn | 1 mol ZnCl ₂ |
| = 41.8 g ZnCl₂ | | | |

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



| | | | |
|--------------------------------|------------|-----------------------|-----------------------|
| 10 g NaCl | 1 mol NaI | 1 mol Cl ₂ | 71 g Cl ₂ |
| | 58.5g NaCl | 2 mol NaCl | 1 mol Cl ₂ |
| = 6.07 g Cl₂ | | | |

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



| | | | |
|------------------------|-------------------------|-------------------------|-----------|
| 4.0g AgNO ₃ | 1 mol AgNO ₃ | 1 mole Cu | 63.5 g Cu |
| | 170g AgNO ₃ | 2 mol AgNO ₃ | 1 mol Cu |

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



| | | |
|------------------------|-----------------------|------------------------|
| 15.1 g Cl ₂ | 1 mol Cl ₂ | 1 mole H ₂ |
| | 71 g Cl ₂ | 1 mole Cl ₂ |

= 0.21 moles H₂

Chemistry Stoichiometry Problems And Answers

Lawrence Wolff



Chemistry Stoichiometry Problems And Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Problems and Solutions for General Chemistry and College Chemistry, Sixth Editions by Nebergall, Holtzclaw, and Robinson Frederick K. Ault, 1980 Foundations of College Chemistry Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis Programmed Problem-solving for First-year Chemistry Arnold B. Loebel, 1983 **Resources in Education** , 1980 Metallurgical & Chemical Engineering Eugene Franz Roeber, Howard Coon Parmelee, 1912 The Pharmacist's Handbook with State Board Examination Questions and Answers International Correspondence Schools, 1923 How to Solve Word Problems in Chemistry David E Goldberg, 2001-06-26 In addition to having to master a vast number of difficult concepts and lab procedures high school chemistry students must also learn with little or no coaching from their teachers how to solve tough word problems Picking up where standard chemistry texts leave off How to Solve Word Problems in Chemistry takes the fear and frustration out of chemistry word problems by providing students with easy to follow procedures for solving problems in everything from radioactive half life to oxidation reduction reactions A Classified Catalogue of Works Published by Longmans, Green & Company , 1912 The United States Catalog Mary Burnham, Carol Hurd, 1928 Questions and answers on the essentials of medical chemistry Lawrence Wolff, 1889 Notes on Quantitative Chemical Analysis Charles William Foulk, 1898 **The essentials of chemical physiology** William Dobinson Halliburton, 1909 The Essentials of Chemical Physiology for the Use of Students William Dobinson Halliburton, 1907 **Journal of Chemical Education** , 1928 **The Popular Science News and Boston Journal of Chemistry** , 1883

Practice Makes Perfect Chemistry Review and Workbook, Second Edition Marian DeWane, Heather Hattori, 2018-12-28
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