

Question 10.10: Stoichiometric calculations

Question 10

The combustion of liquid propan-2-ol (C₃H₈O) proceeds as follows:



(a) Mass A + Mass B

What mass of oxygen is necessary to completely burn 20 g of liquid propan-2-ol?



(b) Mass A + Mass B

What mass of carbon dioxide is formed if 10 g of oxygen is burned?



(c) Mass A + Mass B

What mass of water is produced from 10 g of oxygen?



(d) Mass A + Mass B

Calculate the mass of liquid propan-2-ol that will give 100 g of carbon dioxide.



(e) Mass A + Mass B

What mass of hydrogen atoms (protons) is contained in 10 g of propan-2-ol? What mass of protons is contained in 10 g of oxygen?



Question 11

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