

Chemistry Review Packet

Name: _____

Date: _____

1. Consider the following reaction:



When 2.0 moles of hydrogen are reacted with 48.0 grams of oxygen, how many grams of water are formed?

- A. 18.0 grams B. 36.0 grams C. 54.0 grams D. 72.0 grams E. 90.0 grams

2. Which of the following 1.0 M aqueous solutions has the lowest pH?

- A. HNO_3 B. $NaCl$ C. $NaNO_2$ D. $NaBr$ E. HNO_2

3. Which of the following substances cannot be broken down into simpler substances by physical or chemical means?

- A. Wood B. Baking soda C. Natural gas D. Neon gas E. Water

4. What is the approximate percent composition by mass of hydrogen in the compound ethanol, C_2H_6O (molecular weight, in grams: 46.10 or 46.08 or 46.0)?

- A. 4.3% B. 12.0% C. 29.0% D. 34.0% E. 20.0%

5. The empirical formula of a certain hydrocarbon is C_2H_4 , and it has a molecular weight equal to 80 g/mol. Based on this information, the molecular formula of the hydrocarbon is

- A. C_4H_8 B. C_6H_{12} C. C_8H_{16} D. $C_{10}H_{20}$ E. $C_{12}H_{24}$

6. Which of the following aqueous solutions *all* have the highest normal boiling point?

- A. 0.20 M $CaCl_2$ B. 0.20 M $NaCl$ C. 0.20 M KNO_3 D. 1.0 M $CaCl_2 \cdot 2H_2O$ E. 0.20 M $MgCl_2$

7. Consider the reaction:



How is this reaction best classified?

- A. Precipitation B. Redox reaction C. Single replacement D. Combustion E. Synthesis

8. The following reaction takes place under constant temperature and pressure:



When 100 L of hydrogen gas reacts completely with excess H_2 again, how many liters of gas would form?

- A. 100 L B. 150 L C. 200 L D. 300 L E. 400 L

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F Rizvi



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