

# Experiment 25

## Calorimetry



Andy Watson

A set of nested coffee cups is a good constant pressure calorimeter.

- To determine the specific heat of a metal
- To determine the enthalpy of neutralization for a strong acid–strong base reaction
- To determine the enthalpy of solution for the dissolution of a salt

### OBJECTIVES

The following techniques are used in the Experimental Procedure:

### TECHNIQUES



Accompanying all chemical and physical changes is a transfer of heat (energy); heat may be either evolved (exothermic) or absorbed (endothermic). A **calorimeter** is the laboratory apparatus that is used to measure the quantity and direction of heat flow accompanying a chemical or physical change. The heat change in chemical reactions is quantitatively expressed as the **enthalpy (or heat) of reaction**,  $\Delta H$ , at constant pressure.  $\Delta H$  values are negative for exothermic reactions and positive for endothermic reactions.

### INTRODUCTION

$\Delta H$  values are often expressed as  $\text{J/mol}$  or  $\text{kJ/mol}$ .

Three quantitative measurements of heat are detailed in this experiment: measurements of the specific heat of a metal, the heat change accompanying an acid–base reaction, and the heat change associated with the dissolution of a salt in water.

The energy (heat, expressed in joules,  $J$ ) required to change the temperature of one gram of a substance by  $1^\circ\text{C}$  is the **specific heat**<sup>1</sup> of that substance:

### Specific Heat of a Metal

$$\text{specific heat} \left( \frac{J}{g \cdot ^\circ\text{C}} \right) = \frac{\text{energy } (J)}{\text{mass } (g) \times \Delta T (^\circ\text{C})} \quad (25.1)$$

or, rearranging for energy,

$$\text{energy } (J) = \text{specific heat} \left( \frac{J}{g \cdot ^\circ\text{C}} \right) \times \text{mass } (g) \times \Delta T (^\circ\text{C}) \quad (25.2)$$

$\Delta T$  is the temperature change of the substance. Although the specific heat of a substance changes slightly with temperature, for our purposes, we assume it is constant over the temperature changes of this experiment.

The specific heat of a metal that does not react with water is determined by (1) heating a measured mass of the metal,  $M$ , to a known (higher) temperature; (2) placing it into a measured amount of water at a known (lower) temperature; and (3) measuring the final equilibrium temperature of the system after the two are combined.

<sup>1</sup>The specific heat of a substance is an intensive property (independent of sample size), as are its melting point, boiling point, density, and so on.

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**JG Myers**



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