

## Chapter 2 - Practice Test #1

7. An equilibrium system shifts left when the
- A. rate of the forward reaction is equal to the rate of the reverse reaction.
  - B. rate of the forward reaction is less than the rate of the reverse reaction.
  - C. rate of the forward reaction is greater than the rate of the reverse reaction.
  - D. rate of the forward reaction and the reverse reaction are constant.
8. A 1.00 L flask contains a gaseous equilibrium system. The addition of reactants to this flask results in a
- A. shift left and a decrease in the concentration of the products.
  - B. shift left and a increase in the concentration of the products.
  - C. shift right and a decrease in the concentration of the products.
  - D. shift right and a increase in the concentration of the products.
9. Consider the following equilibrium:  $\text{CH}_4(g) + \text{H}_2\text{O}(g) + \text{heat} = \text{CO}(g) + 3\text{H}_2(g)$   
In which of the following will both stresses shift the equilibrium to the right?
- A. a decrease in temperature and a decrease in volume
  - B. a increase in temperature and a decrease in volume
  - C. a decrease in temperature and a increase in volume
  - D. a increase in temperature and a increase in volume
10. Consider the following equilibrium:  $2\text{SO}_2(g) + \text{O}_2(g) = 2\text{SO}_3(g)$   $\Delta H = -198 \text{ kJ}$   
There will be no shift in this equilibrium when
- A. more  $\text{O}_2$  is added.
  - B. a catalyst is added.
  - C. the volume is increased.
  - D. the temperature is increased.
11. Consider the following equilibrium:  $2\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{O}(g) = \text{Fe}_2\text{O}_3 + 3\text{H}_2$   
The equilibrium expression is
- A.  $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][\text{H}_2]^3}{[\text{Fe}]^2[\text{H}_2\text{O}]^3}$
  - B.  $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][3\text{H}_2]}{[2\text{Fe}][3\text{H}_2\text{O}]}$
  - C.  $\text{Keq} = \frac{[\text{H}_2]^3}{[\text{H}_2\text{O}]^3}$
  - D.  $\text{Keq} = [\text{H}_2]^3$
12. Consider the following equilibrium:  $\text{N}_2\text{O}_4 = 2\text{NO}_2$   $\text{Keq} = 0.133$   
At equilibrium, the  $[\text{N}_2\text{O}_4]$  is equal to

# Chemistry Equilibrium Practice Test

**Burns A. Ralph, Charles Baker, Claire  
Baker**



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**Partial Solutions Manual** Darrell D. Ebbing, David Bookin, 1998

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### **Table of Contents Chemistry Equilibrium Practice Test**

1. Understanding the eBook Chemistry Equilibrium Practice Test
  - The Rise of Digital Reading Chemistry Equilibrium Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Equilibrium Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Equilibrium Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Equilibrium Practice Test

- Personalized Recommendations
  - Chemistry Equilibrium Practice Test User Reviews and Ratings
  - Chemistry Equilibrium Practice Test and Bestseller Lists
5. Accessing Chemistry Equilibrium Practice Test Free and Paid eBooks
    - Chemistry Equilibrium Practice Test Public Domain eBooks
    - Chemistry Equilibrium Practice Test eBook Subscription Services
    - Chemistry Equilibrium Practice Test Budget-Friendly Options
  6. Navigating Chemistry Equilibrium Practice Test eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Equilibrium Practice Test Compatibility with Devices
    - Chemistry Equilibrium Practice Test Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Equilibrium Practice Test
    - Highlighting and Note-Taking Chemistry Equilibrium Practice Test
    - Interactive Elements Chemistry Equilibrium Practice Test
  8. Staying Engaged with Chemistry Equilibrium Practice Test
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Equilibrium Practice Test
  9. Balancing eBooks and Physical Books Chemistry Equilibrium Practice Test
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Equilibrium Practice Test
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Equilibrium Practice Test
    - Setting Reading Goals Chemistry Equilibrium Practice Test
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Equilibrium Practice Test

- 
- Fact-Checking eBook Content of Chemistry Equilibrium Practice Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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