

Chemistry EOC Review – Answer Key

- 1) a. 68,200,000 cg b. 5000 mm c. 0.0125 kL d. 0.548 dm
- 2) -21.5 mL
- 3) production of heat & light, formation of a gas, formation of a precipitate, color change
- 4) $33.64 \text{ g} / 4.28 \text{ ml} = 7.86 \text{ g/ml} = \text{Fe}$
- 5) m.p. $-183^\circ\text{C} = \text{CH}_4$ (methane)
- 6) H_2 – splint test (burning splint ignites hydrogen)
 O_2 – splint test (oxygen reignites glowing splint)
 H_2O – cobalt chloride paper (paper changes from blue to pink in the presence of water vapor)
 CO_2 – splint test (carbon dioxide extinguishes burning splint);
 limewater test (limewater turns cloudy when carbon dioxide is bubbled through it, producing CaCO_3)
- 7) The protons are positively charged particles found inside the nucleus along with the neutrons, which have no charge. The total number of these nucleons in an atom is the mass # / atomic mass of the atom. The electrons are the negatively charged particles found orbiting in "clouds" around the nucleus. Electrons can be gained by the atom to form anions or lost to form cations.
- 8) The thing that makes any element on the periodic table different from all other elements is the number of protons in its nucleus. Isotopes are atoms of the same type of element which have different masses because they have different numbers of neutrons.
- 9) a. $^{39}_{19}\text{K}^+$
 b. $^{57}_{26}\text{Fe}$
 c. $^{32}_{16}\text{S}^{2-}$
- 10)

| <u>chemical symbol</u> | <u>atomic number</u> | <u>mass number</u> | <u># of protons</u> | <u># of electrons</u> | <u># of neutrons</u> | <u>net charge</u> |
|------------------------|----------------------|--------------------|---------------------|-----------------------|----------------------|-------------------|
| K^+ | 19 | 42 | 19 | 18 | 23 | +1 |
| Sr | 38 | 88 | 38 | 38 | 50 | 0 |
| F^- | 9 | 20 | 9 | 10 | 11 | -1 |
| P | 15 | 31 | 15 | 15 | 16 | 0 |
| Ca^{2+} | 20 | 42 | 20 | 18 | 22 | +2 |

- 11) uranium – 238 $238 - 92 = 146$ neutrons
- 12) (1) 199.59g
 (2) They contain the same # of protons (they are the same element), but different #s of neutrons (electrons don't matter, b/c they have virtually no mass), & thus different masses.

Chemistry Eoc Review Answer Key Form 1

Peter J. Holmes, Karin H. Kelder



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Table of Contents Chemistry Eoc Review Answer Key Form 1

1. Understanding the eBook Chemistry Eoc Review Answer Key Form 1
 - The Rise of Digital Reading Chemistry Eoc Review Answer Key Form 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Eoc Review Answer Key Form 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Eoc Review Answer Key Form 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Eoc Review Answer Key Form 1
 - Personalized Recommendations
 - Chemistry Eoc Review Answer Key Form 1 User Reviews and Ratings
 - Chemistry Eoc Review Answer Key Form 1 and Bestseller Lists
5. Accessing Chemistry Eoc Review Answer Key Form 1 Free and Paid eBooks
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 - Chemistry Eoc Review Answer Key Form 1 Budget-Friendly Options
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 - Chemistry Eoc Review Answer Key Form 1 Compatibility with Devices
 - Chemistry Eoc Review Answer Key Form 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Eoc Review Answer Key Form 1
 - Highlighting and Note-Taking Chemistry Eoc Review Answer Key Form 1
 - Interactive Elements Chemistry Eoc Review Answer Key Form 1
8. Staying Engaged with Chemistry Eoc Review Answer Key Form 1
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Eoc Review Answer Key Form 1
9. Balancing eBooks and Physical Books Chemistry Eoc Review Answer Key Form 1
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Eoc Review Answer Key Form 1
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Eoc Review Answer Key Form 1
- Setting Reading Goals Chemistry Eoc Review Answer Key Form 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Eoc Review Answer Key Form 1
- Fact-Checking eBook Content of Chemistry Eoc Review Answer Key Form 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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