

Chemical Quantities II Worksheet A.K.A Stoichiometry

Name: _____ Period: _____ Date: _____

1. Consider the reaction represented by the unbalanced equation



Balance the equation and determine for every 1.00 mol of NH_3 that reacts, how many mol of O_2 is required?

2. A bookcase is to be built from 3 shelves (Sh), 2 side boards (Sb), 1 top (T), 1 base (B), and 4 legs (L). Write a "balanced equation" for the construction of this bookcase.

3. Fill in the missing boxes for the following reaction.

Reactant	+ Reactant	= Product
<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> $\text{N}_2(g)$	$3\text{H}_2(g)$	$2\text{NH}_3(g)$
atoms N_2	6 atoms H_2	<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> molecules NH_3
1 molecule N_2	<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> molecules H_2	<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> molecules NH_3
<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> 6.02×10^{23} molecules N_2	$3 \times 6.02 \times 10^{23}$ molecules H_2	<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> 6.02×10^{23} molecules NH_3
1 mol N_2	<div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> mol H_2	2 mol NH_3
28 g N_2	$3 \times$ <div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> g H_2	$2 \times$ <div style="border: 1px solid black; width: 40px; height: 20px; display: inline-block;"></div> g NH_3

Chemistry Chemical Quantities Section Worksheet

Robert J. Artz



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Holt Chemistry Ralph Thomas Myers, 2004 Chemical

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