



# CHEMISTRY CHEAT SHEET



# EQUILIBRIUM

## EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol  $K_c$ .

**What is Equilibrium?**  
 A reversible reaction is one that can proceed in either direction. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. The concentrations of the reactants and products remain constant over time.

**Equilibrium Constant  $K_c$**   
 The equilibrium constant  $K_c$  is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of their stoichiometric coefficient. It is a constant for a given reaction at a given temperature.



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Reaction	Equilibrium Constant $K_c$
$aA + bB \rightleftharpoons cC + dD$	$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
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Factor	Effect on Equilibrium	Effect on $K_c$
Concentration	<ul style="list-style-type: none"> <li>Increasing the concentration of a reactant or product will shift the equilibrium position to favor the formation of the products or reactants, respectively.</li> </ul>	<ul style="list-style-type: none"> <li><math>K_c</math> is not affected by changes in concentration.</li> </ul>
Pressure	<ul style="list-style-type: none"> <li>Increasing the pressure will shift the equilibrium position to favor the formation of the products if there are fewer moles of gas on the product side.</li> </ul>	<ul style="list-style-type: none"> <li><math>K_c</math> is not affected by changes in pressure.</li> </ul>
Temperature	<ul style="list-style-type: none"> <li>Increasing the temperature will shift the equilibrium position to favor the formation of the products if the reaction is endothermic.</li> </ul>	<ul style="list-style-type: none"> <li><math>K_c</math> is affected by changes in temperature.</li> </ul>

# Chemistry Chemical Equilibrium Study Guide Answers

**B Lingard**



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