



CHEMISTRY CHEAT SHEET



EQUILIBRIUM

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Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c .

What is Equilibrium?		Equilibrium Constant K_c	Graphical Representation of Equilibrium
<p>What is Equilibrium? Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p>	<p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p>	<p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>Equilibrium Constant K_c The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p>	
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Le Chatelier's Principle: the effect of changes in conditions on equilibrium

Change	Effect on Equilibrium	Effect on K_c
Concentration	<ul style="list-style-type: none"> Increasing the concentration of a reactant or product will shift the equilibrium position towards the other side. 	<p>K_c is not affected by changes in concentration.</p>
Pressure	<ul style="list-style-type: none"> Increasing the pressure will shift the equilibrium position towards the side with fewer moles of gas. 	<p>K_c is not affected by changes in pressure.</p>
Temperature	<ul style="list-style-type: none"> Increasing the temperature will shift the equilibrium position towards the endothermic reaction. 	<p>K_c is affected by changes in temperature. For an endothermic reaction, K_c increases with temperature. For an exothermic reaction, K_c decreases with temperature.</p>

Chemistry Chemical Equilibrium Answers Study Guide

Joan Bouillon



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