

## Chapter 5 Review Worksheets Answers

### Section 1:

- C
- D
- B
- C
- D
- C
- Fluorine isotope with mass number 19 and atomic number 9:
  - 9
  - 10
  - $^{19}\text{F}$
- Samarium (Sm) is a lanthanide, what is the element below Sm on the periodic table? Difference in atomic #s?
  - Plutonium Pu
  - 32
- An isotope contains 53 protons, 78 neutrons, & 54 electrons
  - 53
  - 131
  - Iodine
  - Fluorine (F), Chlorine (Cl), Bromine (Br), & Astatine (At)
- Group = Vertical column      Period = Horizontal row
- Atomic mass is the weighted average mass of all the naturally occurring isotopes of the element.  
Atomic number is number of protons found in the nucleus of each atom of the element.
- Ar & K; Co & Ni; Th & Pa; U & Np; Pu & Am; Sg & Bh

### Section 2:

- Identify the element & write its noble gas notation electron configuration:
  - Germanium; Ge [Ar]  $4s^2 3d^{10} 4p^2$
  - Bismuth; Bi [Xe]  $6s^2 4f^{14} 5d^{10} 6p^3$
  - Scandium; Sc [Ar]  $4s^2 3d^1$
  - Radium; Ra [Rn]  $7s^2$
- Identifying off the periodic table:
  - p-block
  - B = Alkali Metals  
C = Alkaline Metals  
D = Transition Metals  
E = Main-Group Elements  
F = Halogens  
G = Noble Gases  
H = Actinides
- Give the symbol, period, group, & block:
  - Sulfur; S, 3<sup>rd</sup> Period, Group 16, p-block
  - Nickel; Ni, 4<sup>th</sup> Period, Group 10, d-block
  - ? [Kr] 5s<sup>1</sup>; Rb, 5<sup>th</sup> Period, Group 1, s-block
  - ? [Ar] 3d<sup>5</sup> 4s<sup>1</sup>; Cr, 4<sup>th</sup> Period, Group 6, d-block      \* Tricky one with an aufbau exception configuration!
- Group #s of blocks:
  - s-block: Group 1 & Group 2
  - p-block: Groups 13-18
  - d-block: Groups 3-12

# Chemistry Chapter Section Review Answers

**RJ Alexander**



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