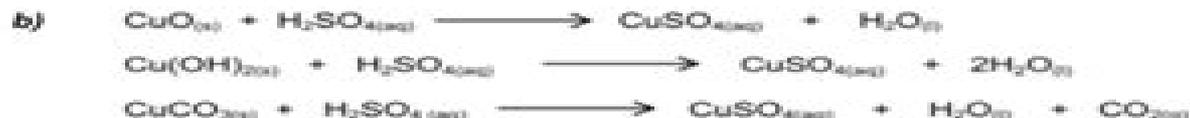


1 a) Copper oxide, hydroxide and carbonate will all react with dilute sulfuric acid. **Copper metal will not react with dilute sulfuric acid.**



2 a) Refer to Fig 9.7, p 73.



b) Hydrogen is extremely flammable. Collect the gas in a test tube. Keep your thumb over the mouth of the test tube. Quickly place the mouth of the test tube in a Bunsen flame, simultaneously removing your thumb. A squeaky 'pop' indicates hydrogen gas.

c) copper, silver, gold

d) calcium, sodium, potassium



(ii) Hydrogen is the least dense element. It rises quickly due to its low density, so the flame (where the hydrogen is) also rises quickly.

- 3 a)
- A = copper
 - B = copper oxide
 - C = copper sulfate
 - D = copper carbonate
 - E = carbon dioxide
 - F = copper nitrate



4 a) The general reaction is *metal + acid* \longrightarrow *salt + hydrogen gas.*

You would test for hydrogen as described in 2 b) above.

The salt formed would be nickel sulfate. You would observe bubbles forming on the metal, the metal would disappear over time and the liquid would become green over time.



c) You would observe the powder fizzing as it reacts with the acid. The gas evolved would be bubbled through limewater to confirm it is carbon dioxide (see Fig 9.7 p 73)



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Paul A. Bartlett, Judith G. Koch



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