

General Chemistry II Chapter 17 Study Guide

Organic acids- carboxylic acids (R-COOH):

- Carboxylic acids (R-COOH): a weak organic acid, contains the carb-oxyl group (-COOH), never fully dissociates in water, an equilibrium process, ex. Formic acid, acetic acid, butyric acid, benzoic acid
 - R= rest of the molecule
- Strength of carboxylic acids: depends on the R group; more electronegative Cl in R group shifts more electron density towards the R group and helps stabilize the structure after the H⁺ ion is removed, thus, chloroacetic acid is the stronger acid in comparison to acetic acid (CH₃COOH)

Organic bases- amines:

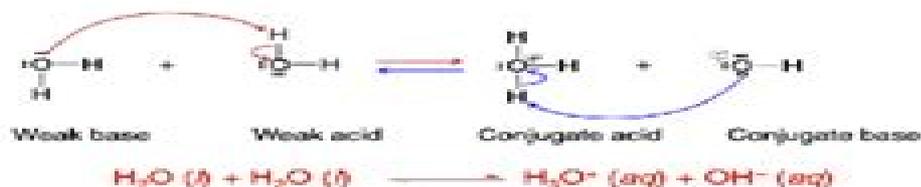
- Amines: organic bases, contains an N atom with a lone pair (ex -NH, -NH₂), ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
 - Only N with lone pair can accept a proton
 - Equilibrium dissociation process
 - NH₄⁺ is an acid and can lose one H⁺ to be neutral

Organic acids and bases (things to keep in mind):

- Acid: H is the only acidic proton; ex. CH₃COOH becomes CH₃COO⁻
- Amine: the proton H⁺ is bonded to N via the lone pair of N; ex. CH₃NH₂ becomes CH₃NH₃⁺
- How about carbon?
 - Protons on C are not acidic, C cannot accept extra protons

Autoionization of water:

- Autoionization: in pure water, one water molecule can donate a proton to another water molecule



At 25 °C, $[\text{H}_3\text{O}^+] = [\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$

$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$
 K_w : ion-product constant

Chemistry Chapter 17 Study Guide Answers

Susan McMurry



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