

$$PV = nRT$$

0.0°C 1 atm
22.414 L

Gas Stoichiometry Practice Sheet

- 1) For the reaction $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(l)}$, how many liters of water can be made from 5 L of oxygen gas and an excess of hydrogen?

$$\frac{5 \cancel{\text{L O}_2}}{22.414 \cancel{\text{L O}_2}} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 10 \text{ L H}_2\text{O}$$

0.0°C
1 atm

- 2) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at STP?

$$\frac{55 \cancel{\text{g O}_2}}{32 \cancel{\text{g O}_2}} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \times 22.414 \text{ L H}_2\text{O} = 77 \text{ L H}_2\text{O}$$

- 3) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at a pressure of 12.4 atm and a temperature of 85°C?

$$\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$$

$$\frac{55 \cancel{\text{g O}_2}}{32.0 \cancel{\text{g O}_2}} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 3.4375 \text{ mol H}_2\text{O}$$

$$(12.4 \text{ atm}) V = (3.4375 \text{ mol H}_2\text{O}) \left(\frac{0.08205 \text{ L atm}}{\text{mol K}} \right) (358 \text{ K})$$

$$V = 8.14 \text{ L H}_2\text{O}$$

- 4) How many liters of water can be made from 34 grams of oxygen gas and

$$\frac{34 \cancel{\text{g O}_2}}{32 \cancel{\text{g O}_2}} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \times 22.414 \text{ L H}_2\text{O} = 47.8 \text{ L H}_2\text{O}$$

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Michael B. McGinnis



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