

## Textbook Questions Unit 1 Biology

### 1.1 - Review Page 18:

1. A) Atom 1 is carbon-13... Atom 2 is carbon-15

B) They're both isotopes of carbon with different amounts of neutrons

2. A) Hydrogen, deuterium, and tritium

B) Radioisotopes are very unstable radioactive isotopes of an element due to unstable combinations of neutrons and protons

3. A)  $1s^2 2s^2 2p^4$

B) Oxygen

4. The repulsion of like negative charges causes separation of bond electron pairs and electron lone electron pairs in a molecule

5. Ionic bonds are the transfer of electrons between oppositely charged ions... while covalent bonds are the sharing of electrons between non-metal atoms

6. Molecular polarity is based on asymmetrical differences in polar dipoles of unequally shared electrons. If the configuration of a molecule with polar covalent bonds is symmetrical, the dipoles would cancel out, meaning that the molecule is nonpolar.

7. These characteristics can affect the solubility, melting/boiling points, and brittleness of molecules.

Stronger Polarity:

- More soluble due to stronger IMF's with  $H_2O$
- Higher MP & BP due to it being harder to break bonds

Larger molecules:

- Less soluble due to it being harder to solute molecules to surround them
- Higher MP & BP due to stronger LDF (more e<sup>-</sup>), and more bonds that take more energy to break to change states

# Biology Review Exercises Unit 18 Answers

**LP Steffe**



## **Biology Review Exercises Unit 18 Answers:**

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for easy identification Canadian content includes Canadian healthcare statistics regarding incidence cultural variations with  
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## **Table of Contents Biology Review Exercises Unit 18 Answers**

1. Understanding the eBook Biology Review Exercises Unit 18 Answers
  - The Rise of Digital Reading Biology Review Exercises Unit 18 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Biology Review Exercises Unit 18 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Biology Review Exercises Unit 18 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Biology Review Exercises Unit 18 Answers
  - Personalized Recommendations
  - Biology Review Exercises Unit 18 Answers User Reviews and Ratings
  - Biology Review Exercises Unit 18 Answers and Bestseller Lists
5. Accessing Biology Review Exercises Unit 18 Answers Free and Paid eBooks
  - Biology Review Exercises Unit 18 Answers Public Domain eBooks
  - Biology Review Exercises Unit 18 Answers eBook Subscription Services
  - Biology Review Exercises Unit 18 Answers Budget-Friendly Options

6. Navigating Biology Review Exercises Unit 18 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Biology Review Exercises Unit 18 Answers Compatibility with Devices
  - Biology Review Exercises Unit 18 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Biology Review Exercises Unit 18 Answers
  - Highlighting and Note-Taking Biology Review Exercises Unit 18 Answers
  - Interactive Elements Biology Review Exercises Unit 18 Answers
8. Staying Engaged with Biology Review Exercises Unit 18 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Biology Review Exercises Unit 18 Answers
9. Balancing eBooks and Physical Books Biology Review Exercises Unit 18 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Biology Review Exercises Unit 18 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Biology Review Exercises Unit 18 Answers
  - Setting Reading Goals Biology Review Exercises Unit 18 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Biology Review Exercises Unit 18 Answers
  - Fact-Checking eBook Content of Biology Review Exercises Unit 18 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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