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What is the percent yield of lead (II) iodide if 48.7 g of KI reacts with excess $\text{Pb}(\text{NO}_3)_2$ to create 48.7 g of product?

$$2\text{KI} + \text{Pb}(\text{NO}_3)_2 \rightarrow 2\text{KNO}_3 + \text{PbI}_2$$

How many moles of sodium will react with water to produce 0.584 mol of hydrogen?

$$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$$

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10 Self Check Activity Chemistry

Zhe Xu



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