

FIRST SEMESTER CHEMISTRY FINAL EXAMINATION FOR GRADE 12TH**CHOOSE THE CORRECT ANSWERS**

- If the solute-solvent interactions are greater than the solute-solute and solvent-solvent interactions, what will be the total vapor pressure of the solution?
A. greater than that calculated from Raoult's law B. less than that calculated from Raoult's law
C. The same as calculated from Raoult's law D. Raoult's law cannot be applied for such interaction
- What is the freezing point of a solution of 250g of CaCl_2 in 1.0kg of water (K_f for water = $1.86^\circ\text{C}/\text{kg}/\text{mol}$) A. -1.3°C B. -6.5°C C. -9.0°C D. -12.55°C
- What is the molarity of a 5 g of hydrogen peroxide in 100mL solution that is used for hair bleaching A. 0.015M B. 0.15M C. 1.5M D. 3M
- What is the osmotic pressure, in atm, of a 0.30M solution of glucose in water that is used for intravenous infusion at body temperature, 37°C ?
A. 0.911 atm B. 773.2atm C. 7.6atm D. 76atm
- Given the unbalanced equation
 $\text{KMnO}_4 + \text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{I}_2 + \text{H}_2\text{O}$ How many grams of KMnO_4 are needed to make 500mL of 0.250N solution? A. 3.95g B. 1.49g C. 1.58g D. 63.2g
- In which of the following cases will the dissolution of sugar be the most rapid?
A. Sugar crystals in hot water B. Sugar crystals in cold water
C. Powdered sugar in hot water D. Powdered sugar in cold water
- Which of the following is not a colligative property
A. Osmotic pressure B. Vapour pressure C. Boiling point D. Freezing point depression
- What is the normality of a 1M solution of H_3PO_4 given the following reaction
 $\text{H}_3\text{PO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{HPO}_4 + \text{H}_2\text{O}$ A. 3N B. 2N C. 6N D. 1N
- A given solution contains 100g of NaCl and 900g of water. what is the mole fraction of NaCl
A. 0.0331 B. 0.9669 C. 0.3 D. 0.7
- A 0.029m aqueous solution of K_2SO_4 has a freezing point of -0.14°C . what are the value of i factors ($K_{f,\text{H}_2\text{O}} = 1.86^\circ\text{C}/\text{m}$) A. 2 B. 4 C. 2.3 D. 2.6
- Which of the following combination never form a buffer solution
A. HF/F^- B. $\text{HCOOH}/\text{HCOO}^-$ C. $\text{HCO}_3^-/\text{CO}_3^{2-}$ D. $\text{HNO}_3/\text{NO}_3^-$
- Which one of the following are true about the graph of weak acid and strong base
A. The PH is equal to 7 at equivalence point B. The graph indicates the titration of strong base with strong acid
C. The PH is greater than 7 at equivalence point D. at half neutralization $\text{PH} = \text{PK}_a$
- A solution in which propionic acid is 0.94% ionized has a PH of 2.85. what is the values of the acid ionization constant (K_a) for propionic acid A. 4.2×10^{-4} B. 6.8×10^{-4} C. 1.3×10^{-5} D. 8.7×10^{-5}
- The acid-base indicator bromocresol green is a weak acid. The yellow acid and blue base forms of the indicator are present in equal concentrations in a solution when the PH is 4.68. what is the PK_a of bromocresol green A. 4.48 B. 4.68 C. 5.68 D. 6.68
- For the step wise dissociation of aqueous H_3PO_4 , which of the following is not conjugate acid-base pair A. HPO_4^{2-} and H_2PO_4^- B. H_3PO_4 and H_2PO_4^- C. H_2PO_4^- and HPO_4^{2-} D. H_2PO_4^- and PO_4^{3-}
- To 0.2M solution of a weak monoprotic acid HA, enough quantity of its sodium salt NaA, was dissolved to give a concentration of 0.2M of the salt. what will be the acid concentration $[\text{H}_3\text{O}^+]$ in the final solution (K_a of HA = 1.8×10^{-5})
A. $1.8 \times 10^{-5}\text{M}$ B. $1.8 \times 10^{-4}\text{M}$ C. 2.8×10^{-5} D. 2.8×10^{-4}
0.0057M B. 0.057M C. 0.57M D. 5.70M
- Consider the following equilibria I. $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$ II. $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NH}_3$,
III. $\text{HSO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{H}_2\text{SO}_3$ In which of the equilibrium does water acts as a Bronsted-lowry base?
A. I only B. II only C. II and III D. I, II, III
- To a 50ml of a buffer solution that is 0.25M in HF and 0.1M in NaF, if 5ml of 0.05M NaOH is added, what is the final PH of this buffer solution, $\text{PK}_a \text{ HF} = 3.17$ A. 2.77 B. 2.80 C. 2.74 D. 3.17
- When sodium acetate is added to an aqueous solution of acetic acid its PH
A. increases B. Decreases C. Remain constant D. may increases or decreases
- How is a buffer solution prepared A. By mixing a weak acid and strong acid B. By mixing a weak acid and its conjugate base C. By mixing a strong acid and its conjugate base D. By mixing strong base and its conjugate acid
- Which one of the followings are not intensive properties A. color B. length C. density D. melting point

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1. Understanding the eBook 100 Question Chemistry Final
 - The Rise of Digital Reading 100 Question Chemistry Final
 - Advantages of eBooks Over Traditional Books
2. Identifying 100 Question Chemistry Final
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 100 Question Chemistry Final
 - User-Friendly Interface
4. Exploring eBook Recommendations from 100 Question Chemistry Final
 - Personalized Recommendations
 - 100 Question Chemistry Final User Reviews and Ratings
 - 100 Question Chemistry Final and Bestseller Lists
5. Accessing 100 Question Chemistry Final Free and Paid eBooks
 - 100 Question Chemistry Final Public Domain eBooks
 - 100 Question Chemistry Final eBook Subscription Services

- 100 Question Chemistry Final Budget-Friendly Options
- 6. Navigating 100 Question Chemistry Final eBook Formats
 - ePub, PDF, MOBI, and More
 - 100 Question Chemistry Final Compatibility with Devices
 - 100 Question Chemistry Final Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 100 Question Chemistry Final
 - Highlighting and Note-Taking 100 Question Chemistry Final
 - Interactive Elements 100 Question Chemistry Final
- 8. Staying Engaged with 100 Question Chemistry Final
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 100 Question Chemistry Final
- 9. Balancing eBooks and Physical Books 100 Question Chemistry Final
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 100 Question Chemistry Final
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 100 Question Chemistry Final
 - Setting Reading Goals 100 Question Chemistry Final
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 100 Question Chemistry Final
 - Fact-Checking eBook Content of 100 Question Chemistry Final
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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