

62. Given the following information:
 Cl_2 bond energy = 239 kJ/mol
 F_2 bond energy = 154 kJ/mol
 $\frac{1}{2}\text{Cl}_2(\text{g}) + \frac{3}{2}\text{F}_2(\text{g}) \rightarrow \text{ClF}_3(\text{g}) \quad \Delta H^\circ = -409 \text{ kJ/mol}$
 calculate the Cl-F bond energy.
- A) 253 kJ/mol
 B) 267 kJ/mol
 C) 136 kJ/mol
 D) 760 kJ/mol

63. Complete the Lewis structure for the molecule:



This molecule has _____ single bonds and _____ multiple bonds.

- A) 6, 3
 B) 11, 5
 C) 11, 2
 D) 13, 0
 E) 4, 2
64. This molecule contains a carbon atom with trigonal planar geometry.
- A) CH_3Cl
 B) C_2H_6
 C) CH_3CHO
 D) CO_2
 E) none of these
65. This molecule is the most polar.
- A) CH_3CHO
 B) CH_3Cl
 C) C_2H_6
 D) CO_2
 E) All are nonpolar.
66. This molecule shows the smallest number of lone pairs in its Lewis structure.
- A) C_2H_6
 B) CH_3Cl
 C) CO_2
 D) CH_3CHO
 E) All have zero lone pairs.

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